

# AP Chem Unit 2

## bonds

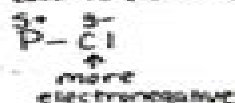
- based on electronegativity



\* electronegativity = measure of how well atoms attract shared electrons

$X-X$   
nonpolar  
(same element)

$X-Y$   
polar  
(diff. element)



\* greater diff. in electronegativity = more polar

\* polar covalent bond = atom w/ greater electronegativity will have partial negative charge

- interstitial alloy: significant difference in atomic radius  
smaller atoms fill in gaps



- substitutional alloy: 2 elements in which the atomic radius is similar



- metal + nonmetal = ionic

- nonmetal + nonmetal = covalent

- good conductor of electricity = metallic bonding

- solid doesn't conduct but as liquid does = ionic bonding

- solid that doesn't conduct & low melting point = covalent molecular

- solid that doesn't conduct & high melting point = covalent network

\* metallic bonding = sea of mobile valence electrons  
↳ good conductors



### Bond Length and Bond Energy

$H-F$

$H-Cl$

$H-Br$

$H-I$

- shortest bond  
- hard to break

- longest bond  
- easy to break

$C-C$

$C=C$

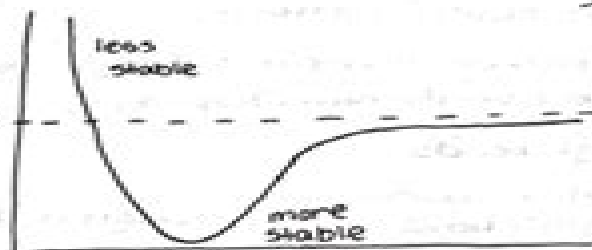
$C \equiv C$

- easy to break  
- longest bond

- hard to break  
- shortest bond

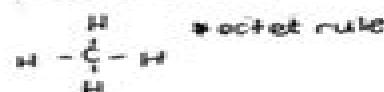
Periodic Trends

as bond energy increases it becomes harder to break and bond length is shorter

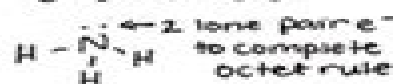


## lewis dot diagrams

$CH_4$  = 8 valence  $e^-$



$NH_3$  = 8 valence  $e^-$



$H_2O$  = 8 valence  $e^-$



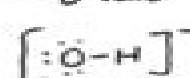
$CH_2Cl_2$  = 20 valence  $e^-$



$NH_4^+$  = 8 valence  $e^-$



$OH^-$  = 8 valence  $e^-$



# Ap Chem Lab 10 Answers

**The Princeton Review**



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