

**A1: Carbon: The Backbone of Life**

- Living organisms are made up of chemicals made mostly out of carbon
- Carbon enters the biosphere through photosynthetic organisms that convert the atoms into the molecules of life
- Carbon has an unparalleled ability to form large, complex, and varied molecules, like proteins, DNA, carbohydrates, etc.
- In organic compounds, carbon is usually bonded to hydrogen, oxygen, nitrogen, phosphorus, or sulfur

**A1: Organic Chemistry is the Study of Carbon Compounds**

- **Organic chemistry** is the study of compounds containing carbon
- Chemists had learned to artificially synthesize inorganic compounds, but it was believed it would be impossible to synthesize organic compounds in a laboratory setting
- Frederick Wöhler accidentally created urea from  $\text{NH}_4^+$  (ammonium) ions and  $\text{CNO}^-$  (cyanate) ions
- Chemists soon learned that compounds that originate in living beings could be reproduced in a laboratory setting
- Organic compounds range from  $\text{CH}_4$  to huge molecules

**A2: Organic Molecules and the Origin of Life on Earth**

- Stanley Miller deduced that complex organic molecules could arise spontaneously under the conditions of early Earth
- Some suggest that spontaneous synthesis of organic compounds created life near a volcano
- The percentages of CHNOPS in each organism are about the same, showing the common ancestry of all species on Earth
- Carbon's versatility in its ability to bond to four different atoms allows an inexhaustible variety of organic molecules to be built
- Species and individuals are differentiated in the types of organic molecules they synthesize, and carbon's versatility contributes to the great biodiversity on Earth

**A1: Carbon Atoms can form Diverse Molecules by Bonding to Four Other Atoms**

- The key to an atom's chemical characteristics is its electron configuration, as it determines how many kinds and number of bonds the atom will form
- The valence shell determines how many electrons are available for bonding

**A2: The Formation of Bonds with Carbon**

- Carbon has six electrons, with two in the first shell and four in the valence shell
- Carbon follows the octet rule of bonding, so it will usually share its four electrons with other atoms
  - A covalent bond is made up of two shared atoms
- Carbon usually forms single or double covalent bond
- Each carbon atom serves as an intersection where a molecule can branch off into four different locations, allowing large and complex molecules to be built
- The hybridization of carbon's four orbitals allows it to form a tetrahedral shape with a  $109.5^\circ$  bond angle
- Any molecule where carbon forms four single bonds has the same tetrahedral shape
- When two carbons are joined via a double bond, each grouping of a carbon to four other atoms has a tetrahedral shape, thus  $\text{C}_2\text{H}_4$  (ethene) is shaped like two overlapping tetrahedrons
- The electron shape and valence electron number of HXNOPS is what serves as the basis for covalent bonding in organic chemistry
- Each line in a structural formula represents one shared pair of electrons

# Ap Biology Chapter 6 Notes

**Yan Bai**



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problem that i foresee is that the configuration of our stage will not allow for a 13 piece orchestra as called for in the score i  
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web the score is full of tricky melodies beautiful ballads and energetic ensemble numbers meaning that there is something  
for everyone if you re looking to learn to sing like one of the cast members of children of eden there are a few things to keep  
in mind first up is warming up your voice

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woodwind 3 woodwind 2 horn u 4 bw 4 clarinet 4 nu w 4 u 4 4 b j 4 4 oboe 1 b b p b b p n p flute b b 4 w w w w p b nw 3 2

solo w cello espr p 4 p w p 5 storytellers s a w wpiano

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