

AP Chemistry Test (Chapter 13)

Multiple Choice (30%)

- 1) Which one best describes the K_c for this reaction?



A) $K_c = \frac{[\text{A}]^3[\text{D}]^2}{[\text{E}]^4[\text{D}]^2}$

B) $K_c = \frac{[\text{B}]^3[\text{E}]^4[\text{G}]^2}{[\text{A}]^3[\text{D}]^2[\text{C}]^5}$

C) $K_c = \frac{[\text{E}]^4[\text{B}]^3}{[\text{A}]^3[\text{D}]^2}$

D) $K_c = \frac{[\text{A}]^3[\text{D}]^2[\text{C}]^5}{[\text{B}]^3[\text{E}]^4[\text{G}]^2}$

E) None of these is correct

- 2) Please consider the gas phase reaction:



Which one would decrease the yield of H_2 ?

A) High T & high P

B) Low T & low P

C) High T & low P

D) Low T & high P

- 3) AlBr_3 dissolves in water, according to this mechanism:



What effect would the addition of Na_3PO_4 have on this equilibrium?

A) Q becomes larger & the reaction creates more products.

B) Q becomes larger & the reaction creates more reactants.

C) Q becomes smaller & the reaction creates more products.

D) Q becomes smaller & the reaction creates more reactants.

E) There is no net effect on the equilibrium position.

- 4) Based upon the following information, which one is true about the reaction mixture?



$$P(\text{NH}_3) = 2.19 \text{ atm} \quad P(\text{N}_2) = 4.33 \text{ atm} \quad P(\text{H}_2) = 0.72 \text{ atm}$$

A) The product concentrations are too high & the reaction will shift left to establish equilibrium.

B) The product concentrations are too high & the reaction will shift right to establish equilibrium.

C) The reactant concentrations are too high & the reaction will shift right to establish equilibrium.

D) The reactant concentrations are too high & the reaction will shift left to establish equilibrium.

Ap Chemistry Chapter 13 Test

**Peter J. Mikulecky, Michelle Rose
Gilman, Kate Brutlag**



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