

Antacid Analysis and Titration – Lab Report Assistant

Exercise 1: Back Titration of Antacid Neutralization

Data Table 1. Antacid Neutralization Data

Mass of Crushed Antacid (g)	
Concentration of HCl (M)	1M
Volume HCl (mL)	5 mL
Concentration of NaOH (M)	1M
Initial NaOH Volume (mL)	
Final NaOH Volume (mL)	
Total Volume of NaOH Used (mL)	

Data Table 2. Experimental Results

HCl available for neutralization (g):	0.3646 (g)
Moles of NaOH required to reach stoichiometric point:	0.0039 (moles)
HCl neutralized by antacid:	0.222 (g)
HCl neutralized per gram of antacid:	0.445 (g)

Questions:

- If an antacid tablet weighed 1.6 grams, how many moles of gastric acid (HCl) would it neutralize? Use the results obtained in **Data Tables 1 and 2** to explain and quantify your answer.
- If you performed this experiment with a titrant of 0.5M NaOH, would you expect your results in **Data Table 2** to change or stay the same? Explain your answer.
- The reaction that occurred when the antacid mixed with the HCl resulted in an additional product (besides a salt and water). Did you see evidence of this product? Describe the experimental evidence you witnessed that supports the formation of the additional product.



Antacid Analysis Lab Answers

Neil D. Jespersen, Pamela Kerrigan

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author names.

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