

CHEM EXAM 3 Fall 2009 - Form A

Fill in your name, CID and section on this test booklet. Fill in and bubble in your name, CID (for the letter "C" bubble zero), section and "001" on the scantron form. For questions #66, mark A. For each question, read all possible choices then mark the best one on your scantron. Use a #2 pencil and make clean answers to insure proper scoring. When finished, turn in your scantron and show this booklet and your CID to the proctor. All questions are worth 4 points each. Let your knowledge show!

1. Which of the following is balanced correctly?

- A. $\text{CH}_4 + 2 \text{O}_2 \rightarrow \text{CO}_2 + 2 \text{H}_2\text{O}$
 B. $6 \text{H}_2\text{O} + \text{P}_4\text{O}_{10} \rightarrow 4 \text{H}_3\text{PO}_4$
 C. $3 \text{H}_2 + 3 \text{N}_2 \rightarrow 2 \text{NH}_3$
 D. $2 \text{C}_2\text{H}_6\text{O} + 3 \text{O}_2 \rightarrow 2 \text{CO}_2 + 4 \text{H}_2\text{O}$

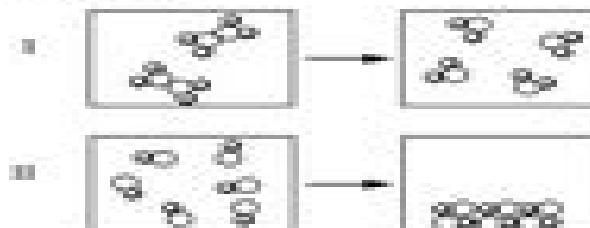
A chemical equation is balanced when the same number of atoms are in the reactants and products. The reactants in B contain 16 oxygen atoms, but the products contain only 12 (B). The reactants in C contain 6 nitrogen atoms, but the product contains only 2 (C). The reactants in D contain 6 carbon atoms, 16 hydrogen atoms, and 8 oxygen atoms. The products contain 2 carbon atoms, 8 oxygen atoms, and 8 hydrogen atoms (D).

2. Which of the following is an oxidation-reduction reaction?

- A. $2 \text{NaCl (aq)} + \text{Ca(NO}_3)_2 \text{ (aq)} \rightarrow 2 \text{NaNO}_3 \text{ (aq)} + \text{CaCl}_2 \text{ (aq)}$
 B. $\text{CrCl}_3 \text{ (aq)} + 3 \text{KOH (aq)} \rightarrow 3 \text{KCl (aq)} + \text{Cr(OH)}_3 \text{ (s)}$
 C. $\text{KOH (aq)} + \text{HCl (aq)} \rightarrow \text{KCl (aq)} + \text{H}_2\text{O (l)}$
 D. $2 \text{CrCl}_3 \text{ (aq)} + \text{CuCl}_2 \text{ (aq)} \rightarrow 2 \text{CrCl}_2 \text{ (aq)} + \text{Cu (s)}$

An oxidation-reduction reaction is one in which electrons are transferred from one or more atoms in the reactants to one or more atoms in the products. The reactants and the products in A are all strong electrolytes (a net ionic equation, A). Two aqueous solutions react in B to form a solid precipitate, B. A strong acid reacts with a strong base in C to form an ionic salt and water (acid-base, C). Two aqueous solutions react to form a solid in D, but chromium in CrCl_3 has an oxidation number of +3. In CrCl_2 , it has an oxidation number of +2. Copper in CuCl_2 has an oxidation number of +2. Solid copper has an oxidation number of 0. Therefore, chromium is oxidized while copper is reduced (an oxidation-reduction reaction).

3. Which of the following transfers heat to the surroundings?



- A. I only
 B. II only
 C. Both I and II
 D. neither I or II

In Figure I, the bond between the two white spheres is broken. In Figure II, six molecules combine to form a solid precipitate. Bond breaking requires energy (endothermic), while bond forming releases energy (exothermic).

4. How many grams of N_2H_4 would be required to react completely with 19.0 g of O_2 according to the balanced chemical equation?



- A. 37.1 g
 B. 42.1 g
 C. 74.2 g
 D. 23.2 g
 E. 26.5 g

The chemical equation is used to relate the number of moles of one substance to the number of moles of another. If grams of O_2 is given, it needs to be converted to moles. The coefficients are used as mole ratios, and the formula mass of N_2H_4 is used to get back to grams:

$$\begin{aligned} 19.0 \text{ g O}_2 & \left(\frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \right) \left(\frac{2 \text{ mol N}_2\text{H}_4}{3 \text{ mol O}_2} \right) \left(\frac{32 \text{ g N}_2\text{H}_4}{1 \text{ mol N}_2\text{H}_4} \right) \\ & = 37.1 \text{ g N}_2\text{H}_4 \end{aligned}$$

Answer Key To Chem 141lab Manuel

Richard A. McPherson



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Conceptual Physical Science Paul G. Hewitt, John Suchocki, Leslie A. Hewitt, 2004 *Conceptual Physical Science* Third Edition takes learning physical science to a new level by combining Hewitt's leading conceptual approach and friendly writing style in a new edition that provides stronger integration of the sciences more quantitative coverage and a wealth of new media resources to help professors in class and students out of class The book's consistent high quality coverage includes five new chapters on chemistry astronomy and earth science for an even more balanced approach to physical science New Looking Forward and Looking Back boxes connect themes and concepts throughout the book helping students see the big picture More computational coverage eg Figuring Physical Science in chapter calculation allows students to practice the quantitative skills they need to master the concepts of physical science and be able to apply their knowledge Looking Forward and Looking Back boxes in every chapter connect themes and concepts throughout the book helping students see the big picture of physical science Powerful media package includes a comprehensive suite of award winning interactive online tutorials that offer students 24/7 help A media gri **Hayes' Principles and Methods of Toxicology** A. Wallace Hayes, Tetyana Kobets, 2023-07-03 Hayes Principles and Methods of Toxicology has long been established as a reliable and informative reference for the concepts methodologies and assessments integral to toxicology The new edition contains updated and new chapters with the addition of new authors while maintaining the same high standards that have made this book a benchmark resource in the field Key Features The comprehensive yet concise coverage of various aspects of fundamental and applied toxicology makes this book a valuable resource for educators students and professionals Questions provided at the end of each chapter allow readers to test their knowledge and understanding of the material covered All chapters have been updated and over 60 new authors have been added to reflect the dynamic nature of toxicological sciences New topics in this edition include Safety Assessment of Cosmetics and Personal Care Products The Importance of the Dose Rate Response Novel Approaches and Alternative Models Epigenetic Toxicology and an Expanded Glossary The volume is divided into 4 major sections addressing fundamental principles of toxicology Section I Principles of Toxicology major classes of established chemical hazards Section II Agents current methods used for the assessment of various endpoints indicative of chemical toxicity Section III Methods as well as toxicology of specific target systems and organs Section IV Organ and System Specific Toxicology This volume will be a valuable tool for the audience that wishes to broaden their understanding of hazards and mechanisms of toxicity and to stay on top of the emerging methods and concepts of the rapidly advancing field of toxicology and risk assessment *Chemistry* McGraw-Hill Staff, 2001-07

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