

## Thermochemistry - Answers

1. Given:  $\text{N}_2(\text{g}) + 3 \text{H}_2(\text{g}) \rightarrow 2 \text{NH}_3(\text{g}) + 92.5 \text{ kJ}$ , what is  $\Delta H_f^\circ$  for  $\text{NH}_3(\text{g})$ ?

$$\begin{aligned}\Delta H_{\text{rxn}} &= (2 \text{ mol})(\text{NH}_3) - (1 \text{ mol})(\text{N}_2) - (3 \text{ mol})(\text{H}_2) \\ -92.5 \text{ kJ} &= (2 \text{ mol})(\text{NH}_3) - (1 \text{ mol})(0 \text{ kJ/mol}) - (3 \text{ mol})(0 \text{ kJ/mol}) \\ -92.5 \text{ kJ} &= (2 \text{ mol})(\text{NH}_3)\end{aligned}$$

$$\Delta H_f \text{ for } \text{NH}_3 = -46.2 \text{ kJ/mol}$$

2. Given:
- |  |                              |
|--|------------------------------|
| $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$   | $\Delta H = 175 \text{ kJ}$  |
| $\text{Ca}(\text{OH})_2(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{CaO}(\text{s})$                               | $\Delta H = 67 \text{ kJ}$   |
| $\text{Ca}(\text{OH})_2(\text{s}) + 2 \text{HCl}(\text{g}) \rightarrow \text{CaCl}_2(\text{s}) + 2 \text{H}_2\text{O}(\text{l})$ | $\Delta H = -198 \text{ kJ}$ |

Calculate  $\Delta H_{\text{rxn}}$  for  $\text{CaCO}_3(\text{s}) + 2 \text{HCl}(\text{g}) \rightarrow \text{CaCl}_2(\text{s}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$

$\text{CaCO}_3(\text{s}) \rightarrow \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$	$\Delta H = (1)(175 \text{ kJ})$
$\text{Ca}(\text{OH})_2(\text{s}) + 2 \text{HCl}(\text{g}) \rightarrow \text{CaCl}_2(\text{s}) + 2 \text{H}_2\text{O}(\text{l})$	$\Delta H = (1)(-198 \text{ kJ})$
<u><math>\text{H}_2\text{O}(\text{l}) + \text{CaO}(\text{s}) \rightarrow \text{Ca}(\text{OH})_2(\text{s})</math></u>	<u><math>\Delta H = (-1)(67 \text{ kJ})</math></u>



3. The heat of combustion of naphthalene,  $\text{C}_{10}\text{H}_8(\text{s})$  is  $-5162 \text{ kJ}$ . What is the heat of formation of naphthalene?



$$\begin{aligned}\Delta H_{\text{comb}} &= (10 \text{ mol})(\text{CO}_2) + (4 \text{ mol})(\text{H}_2\text{O}) - (1 \text{ mol})(\text{C}_{10}\text{H}_8) - (12 \text{ mol})(\text{O}_2) \\ -5162 \text{ kJ} &= (10 \text{ mol})(-393.5 \text{ kJ/mol}) + (4 \text{ mol})(-285.8 \text{ kJ/mol}) - (1 \text{ mol})(\text{C}_{10}\text{H}_8) - (12 \text{ mol})(0 \text{ kJ/mol}) \\ -5162 \text{ kJ} &= -3935.0 \text{ kJ} - 1143.2 \text{ kJ} - (1 \text{ mol})(\text{C}_{10}\text{H}_8) - 0 \text{ kJ} \\ -5162 \text{ kJ} + 1143.2 \text{ kJ} + 3935.0 \text{ kJ} &= -(1 \text{ mol})(\text{C}_{10}\text{H}_8) \\ -83.8 \text{ kJ} &= -(1 \text{ mol})(\text{C}_{10}\text{H}_8)\end{aligned}$$

$$\Delta H_f \text{ for } \text{C}_{10}\text{H}_8 = +83.8 \text{ kJ/mol}$$

# Answers To Thermochemistry

**Peter J. Mikulecky, Chris Hren**



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**General Chemistry Workbook** Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material

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**A Class Book of Physical Chemistry** Thomas Martin Lowry, Samuel Sugden, 1929

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