

CH111 Nomenclature Worksheet

Provide names/formulas for the following substances as requested.
Also indicate whether the substance is expected to be ionic or covalent in its bonding.

Compound: name	Compound Formula	Bonding (ionic, covalent, both)
1. potassium bromide	<u>KBr</u>	<u>ionic</u>
2. <u>lithium oxide</u>	<u>Li₂O</u>	<u>ionic</u>
3. sodium nitride	<u>Na₃N</u>	<u>ionic</u>
4. calcium chloride	<u>CaCl₂</u>	<u>ionic</u>
5. <u>calcium sulfide</u>	<u>CaS</u>	<u>ionic</u>
6. aluminum iodide	<u>AlI₃</u>	<u>ionic</u>
7. aluminum phosphide	<u>AlP</u>	<u>ionic</u>
8. sulfur trioxide	<u>SO₃</u>	<u>covalent</u>
9. <u>diphosphorus pentoxide</u>	<u>P₂O₅</u>	<u>covalent</u>
10. <u>nitrogen dioxide</u>	<u>NO₂</u>	<u>covalent</u>
11. <u>iron(III) chloride</u>	<u>FeCl₃</u>	<u>ionic</u>
12. dinitrogen pentoxide	<u>N₂O₅</u>	<u>covalent</u>
13. gallium carbonate	<u>Ga₂(CO₃)₃</u>	<u>both</u>
14. hydrobromic acid	<u>HBr</u>	<u>ionic</u>
15. tin(IV) acetate	<u>Sn(C₂H₃O₂)₄</u>	<u>both</u>
16. <u>sulfuric acid</u>	<u>H₂SO₄</u>	<u>both</u>
17. copper(II) sulfate	<u>CuSO₄</u>	<u>both</u>
18. <u>iron(II) nitrate</u>	<u>Fe(NO₃)₂</u>	<u>both</u>
19. calcium nitrite	<u>Ca(NO₂)₂</u>	<u>both</u>
20. <u>sulfuric acid</u>	<u>H₂SO₄</u>	<u>both</u>

(Yes, 20 is a duplicate of 16.)

NOTE: ANYTIME a metal is present, the bonding HAS **IONIC CHARACTER**. If the anion ("-" ion) is polyatomic, then **BOTH IONIC and COVALENT** bonding are present.

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