

Autolysis is when one of the products of the reaction is the catalyst for the reaction.

- Starts slowly (no catalysts) $\rightarrow$ speeds up (catalysts form) $\rightarrow$ slows down (lack of reactant).



Acid-base nature of Aqua Ions

M^{3+} is more acidic than M^{2+} since they are smaller & more highly charged $\rightarrow$ more strongly polarising $\rightarrow$ strongly attracts the water ligands, weakening the O-H bond, releasing a proton into the solution $\rightarrow$ more acidic.

- Carbonates of M^{3+} do not exist – M^{3+} is more acidic $\rightarrow$ carbonate ions can remove protons from $[\text{M}(\text{H}_2\text{O})_6]^{3+}$ to form $[\text{M}(\text{H}_2\text{O})_5(\text{OH})_2]$, but cannot do so from $[\text{M}(\text{H}_2\text{O})_6]^{2+} \rightarrow \text{M}^{2+}$ forms MCO_3 & M^{3+} forms CO_2 & its hydroxide.

M^{2+} acid-base reactions



M^{3+} acid-base reactions



Ligand Substitution Reactions

When replacing water as a ligand – the replacement can be complete / partial.



NH_3 can act as a base before as a ligand $\therefore$ redissolves



A2 Empa Notes Chemistry

Michael Rajnik

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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