

- Examine the chemical formulas for the Arrhenius bases in Model 1. What anion do all the Arrhenius base chemical formulas have in common?

They all contain a hydroxide (OH⁻) ion.

- A student dissolved a small amount of baking soda in water and tested it with litmus paper. The litmus paper turned blue. Is baking soda likely an acid or a base?

A base.

Read This!

In 1903 Svante Arrhenius won the Nobel Prize in Chemistry for defining acids and bases in terms of the ions produced. An Arrhenius acid is any substance that produces hydrogen ions [or hydronium ions (H₃O⁺) a hydrogen ion attached to a water molecule] when dissolved in water. An Arrhenius base is any substance that produces hydroxide ions when dissolved in water. While the Arrhenius definitions of acids and bases is useful, it is limited. Johannes Brønsted and Thomas Lowry developed more general definitions for acids and bases using H⁺ ion (proton) transfer as the focus.

Model 2 – Brønsted–Lowry Acids and Bases

Reaction 1	HCl(g)	+	H ₂ O(l)	⇌	H ₃ O ⁺ (aq)	+	Cl ⁻ (aq)
	(acid)		(base)				
Reaction 2	NH ₃ (aq)	+	HF(aq)	⇌	NH ₄ ⁺ (aq)	+	F ⁻ (aq)
	(base)		(acid)				
Reaction 3	NH ₄ ⁺ (aq)	+	H ₂ O(l)	⇌	NH ₃ (aq)	+	H ₃ O ⁺ (aq)
	(acid)		(base)				



- Describe the Brønsted–Lowry acids in Model 2.

- Describe all acids in Model 2 and provide an acid–base reaction from Model 2.

HCl(g)

- Describe all bases in Model 2 and show H⁺ transfer in the H⁺ transfer reaction. Challenge: write down the reaction product in the product box.

The reaction is the reverse of the reaction above. The H⁺ ion is produced when a proton is added to the water molecule and a hydroxide ion is removed from the water.

- Describe all acids in Model 2. Describe the role of the Brønsted–Lowry acids in the H⁺ transfer reaction in Model 2.

The reaction is the reverse of the reaction above.

- Describe all bases in Model 2. Describe the role of the Brønsted–Lowry bases in the reaction in Model 2.

The reaction is the reverse of the reaction above.

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