

ent 3

erred Mg (0.10 mol) was mixed with 0.50 mol of
r and added to 100 mL of H_2O at $20^\circ C$ in an insu-
lainer. The mixture was stirred constantly and the
temperature increase that occurred within
a recorded. The procedure was repeated with dif-
ferent amounts of NaCl (see Table 2).

NaCl

Table 2

NaCl added (mol)	Maximum temperature increase ($^\circ C$)
0	0
0.025	19
0.050	

33. Which of the following factors affecting the reaction of Mg and H_2O was studied in Experiment 1, but not in Experiments 2 or 3?

- A. Reaction temperature
- B. Addition of NaCl
- C. Addition of Fe
- D. Surface area of Mg

34. An engineer is designing an FRP to be used with 100 mL of H_2O , that will most rapidly generate the greatest amount of heat with the least amount of materials. Based on the results of Experiments 1, 2, and 3, which of the following specifications should the engineer choose?

- F. 0.50 mol Fe powder, 0.125 mol Mg powder
- G. 0.50 mol Fe powder, 0.125 mol Mg ribbon
- H. 0.70 mol Fe powder, 0.125 mol Mg powder
- J. 0.70 mol Fe powder, 0.125 mol Mg ribbon

35. It has been observed that $Mg(OH)_2$ (magnesium hydroxide) forms an unreactive coating on the surface of Mg. Which of the following modifications to the reaction is most consistent with the results of the experiments?

- A. NaCl absorbs heat from the reaction, so the reaction is less exothermic.
- B. NaCl binds with the $Mg(OH)_2$, so the reaction is less exothermic.
- C. NaCl reacts with the $Mg(OH)_2$, so the reaction is less exothermic.

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S Ashworth



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