

ACS Practice Test 1

Acids & Bases:

- All are examples of Lewis acid–base reactions *except*
 - $\text{Cu}^{2+}_{(aq)} + 4\text{NH}_3_{(aq)} \rightleftharpoons [\text{Cu}(\text{NH}_3)_4]^{2+}_{(aq)}$
 - $\text{HCl}_{(g)} + \text{NH}_3_{(g)} \rightarrow \text{NH}_4\text{Cl}_{(s)}$
 - $\text{H}^+_{(aq)} + \text{OH}^-_{(aq)} \rightleftharpoons \text{H}_2\text{O}_{(l)}$
 - $2\text{Na}_{(s)} + \text{Cl}_{2(g)} \rightarrow 2\text{NaCl}_{(s)}$
- According to the Lewis definition, an acid is a species
 - having a hydrogen ion.
 - donating a pair of electrons.
 - accepting a pair of electrons.
 - accepting a hydrogen ion.
- All are potential Lewis bases *except*
 - NH_3
 - H_2O
 - CH_4
 - CN^-
- Which hydrated metal ion is most acidic under conditions of equal molar concentration in water?
 - Al^{3+}
 - Ba^{2+}
 - K^+
 - Zn^{2+}
 - Ag^+
- Which element in Group 5A forms the most basic oxide?
 - N
 - P
 - As
 - Sb
 - Bi

Atomic Structure:

- In what respect does an atom of magnesium differ from a magnesium ion (Mg^{2+})?
 - The ion has a more stable electronic arrangement than the atom.
 - The positive charge on the nucleus of the ion is two units greater than the nuclear charge on the atom.
 - The ion has two more protons than the atom.
 - The ion will react more readily with free Cl_2 .
 - The ion has two more planetary electrons than the atom.
- A certain atom has the symbol **Px**. What does this symbol tell about an atom of the element?
 - It has 148 electrons.
 - It has a valence of 4.
 - It has 221 protons in the nucleus.
 - It has 75 neutrons in the nucleus.
 - Its atomic mass is 73 u.
- The atomic mass of an element is 32.07 u and its atomic number is 16. The number of protons contained in the nucleus of an atom of this element is
 - 16
 - 31
 - 32
 - 48
- The atomic mass of an element is 32.07 u and its atomic number is 16. The element forms a simple ion. The ion will most likely have a charge of
 - 1–
 - 2–
 - 3+
 - 1+

Acs Test Practice 1

Perry J. Pickhardt, David H. Kim



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