

Combustion

Complete: $\text{---} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

Incomplete: $\text{---} + \text{O}_2 \rightarrow \text{CO} + \text{H}_2\text{O}$

$\text{---} + \text{O}_2 \rightarrow \text{C} + \text{H}_2\text{O}$

Alkane/Ether

Balance Carbon, Hydrogen, Oxygen

Halogenation

Alkene: $\text{H}_2\text{C}=\text{CH}_2 + \text{Br}_2 \rightarrow \text{H}_2\text{C}(\text{Br})\text{CH}_2(\text{Br})$

Alkyne: same as alkene, but double addition reagent

$\rightarrow \text{Cl}_2$ or Br_2

Addition of Simple Acids

Alkene: $\text{CH}_2=\text{CH}_2 + \text{HCl} \rightarrow \text{CH}_3\text{CH}_2(\text{Cl})$

Alkyne: Same as Alkene, but twice addition reagent

$\rightarrow \text{HBr}$ or HCl

Markovnikov's Rule: Hydrogen goes to Carbon with more Hydrogen; the rich get richer

Hydration

Alkene: $\text{H}_2\text{C}=\text{CH}_2 + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{CH}_2(\text{OH})$

Alkyne: Same as Alkene, but twice the addition reagent

$\rightarrow \text{H}_2\text{O}$ & Acid Catalyst (PL: Pd, Ni)

Markovnikov's Rule: the rich get richer

Hydrogenation

Alkene: $\text{H}_2\text{C}=\text{CH}_2 + \text{H}_2 \rightarrow \text{CH}_3\text{CH}_3$

Alkyne: Same as Alkene, but twice the addition reagent

Aldehyde: $\text{CH}_2(\text{O}) + \text{H}_2 \rightarrow \text{CH}_2(\text{OH})$

Ketone: $\text{H}_3\text{CC}(\text{O})\text{CH}_3 + \text{H}_2 \rightarrow \text{H}_3\text{CCH}_2\text{CH}_3$

Alkene/Alkyne \rightarrow Alkane
Aldehyde \rightarrow Primary Alcohol
Ketone \rightarrow Secondary Alcohol

Polymerization

Addition: $\text{CH}_2=\text{CH}=\text{CH}=\text{CH}_2 \rightarrow \text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2$

Condensation: $(\text{OH})\text{C}(\text{O})\text{C}(\text{OH}) + \text{HOCH}_2\text{OH} \rightarrow$

Polyamides: $(\text{O})\text{C}(\text{O})\text{C}(\text{O})\text{C}(\text{O}) + \text{H}_2\text{N-NH}_2 \rightarrow$

Addition Polymerization: linking together many Alkene molecules through addition reactions

Carboxylic Acid: Condensation

Amide: Polyamide

Substitution

Arom: Switch 1 Hydrogen with one of the H 's addition reagents

Aromatics: stable/chemically inert

Dehydration

180°C: $\text{H}_3\text{CCH}_2(\text{OH}) \rightarrow \text{H}_2\text{C}=\text{CH}_2 + \text{H}_2\text{O}$

140°C: $\text{H}_3\text{CCH}_2\text{OH} + \text{H}_3\text{CCH}_2\text{OH} \rightarrow \text{H}_3\text{CCH}_2\text{OCH}_3 + \text{H}_2\text{O}$

@180°C Alcohol \rightarrow Alkene

@140°C Alcohol + Alcohol \rightarrow Ether

Requires Acid Catalyst (H_2SO_4)

Oxidation

Primary Alcohol \rightarrow Aldehyde

Secondary Alcohol \rightarrow Ketone

Tertiary Alcohol \rightarrow NR

Aldehyde \rightarrow Carboxylic Acid

Ketone \rightarrow NR

Thiols \rightarrow Disulfide

Oxidizing Agents: $\text{K}_2\text{Cr}_2\text{O}_7$, KMnO_4

Tollens Reagent: $2\text{Ag}(\text{NH}_3)_2^+ \rightarrow 2\text{Ag}$

Benedict's Reagent: $2\text{Cu}^{2+} \rightarrow \text{Cu}_2\text{O}$ (blue) \rightarrow (red precipitate)

Dissection

Acid: Carboxylic Acid, Phenol

Donate H from $-\text{OH}$ to $\text{H}_2\text{O} \rightarrow \text{O}^- + \text{H}_3\text{O}^+$

Base: Amine

Steal H from $\text{H}_2\text{O} \rightarrow \text{NH}_4^+ + \text{OH}^-$

Acid + $\text{H}_2\text{O} \rightarrow$

Base + $\text{H}_2\text{O} \rightarrow$

Neutralization

Acid + Base \rightarrow Salt + H_2O

Salt Ex: OK^+

Base: KOH , NaOH

Acid: Phenol, Carboxylic Acid



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Matthew Bowling



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