ACS Practice Test 1

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(B)
$$HCl_{00} + NH_{100} \rightarrow NH_{4}Cl_{00}$$

(C)
$$H^*_{(eq)} + OH^*_{(eq)} \Rightarrow H_2O_6$$

(D)
$$2Na_{\theta l} + Cl_{2\theta l} \rightarrow 2NaCl_{\theta l}$$

All are notential Lewis bases except

4. Which hydrated metal ion is most acidic under conditions of equal molar concentration in water?

5. Which element in Group 5A forms the most basic oxide?

Atomic Structure:

6 In what respect does an atom of magnesium differ from a magnesium ion (Mg2*)?

(A)The ion has a more stable electronic arrangement than the atom.

(B) The positive charge on the nucleus of the ion is two units greater than the nuclear charge on the atom.

(C)The ion has two more protons than the atom.

(D)The ion will react more readily with free Cl.

(E)The ion has two more planetary electrons than the atom.

7. A certain atom has the symbol Px. What does this symbol tell about an atom of the element?

- It has 148 electrons CAD
- It has a valence of 4.
- (C)It has 221 protons in the nucleus.
- (ID) It has 75 neutrons in the nucleus

Its atomic mass is 73 u. (E)

3. The atomic mass of an element is 32.07 u and its atomic number is 16. The number of protons contained in the nucleus of an atom of this element is

- 3+
- 5 +

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Klaus-Martin Goeters

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