
Inorganic Chemistry ACS: Study Guide Exam|334 Questions with Accurate Answers

Ionization energy - -energy required to remove the least tightly bound electron from a neutral atom in the gas phase

- periodic trend of ionization energy - -highest at top right-smaller electron=harder to remove

- Why is a half filled subshell so stable? - -it serves to maximize the stabilizing interactions while minimizing the destabilizing interactions among electrons

- exchange interaction - -pie, stabilizing, result of electrons pairing in degenerate orbitals with parallel spin

- pairing energy - -destabilizing, coulomb interaction, pic, energy of electron-electron repulsion in a filled orbital

- Is it easier to ionize a high energy or low energy electrons - -high energy electron-already contains more energy so it requires less energy input

- What happens when a 3d series metal is ionized? - -the first electron to be ionized will come from the 4s orbital, the other s electron will enter the d orbital ($4s^0 3d^{n+1}$)

- lanthanide contraction - -reduction in atomic radius following the lanthanide series, contrary to the overall trend observed for the periodic table

- lanthanides - -elements 57-71, first appearance of f orbitals, f orbitals are poor at shielding so any electrons added will have a higher Z_{eff} , shrinking the radius

- Slater's rules - -tell us what the effective nuclear charge will be, $Z_{eff} = Z - \sigma$, Z is the atomic number, σ =sum of the number of electrons in a given subshell multiplied by a weighting coefficient (page 1)

- Shielding - -the reduction in charge attraction between the nucleus and electrons due to electrons between the nucleus and the electron in question, it is considered the be between if it has a lower energy

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Section three elaborates on the synthetic chemistry of five important classes of inorganic functional materials namely ordered porous materials carbon materials advanced ceramic materials host guest materials and hierarchically structured materials Section four consists of four chapters where the synthesis of functional inorganic aggregates is discussed giving special attention to the growth of single crystals assembly of nanomaterials and preparation of amorphous materials and membranes The new edition s biggest highlight is Section five where the frontier in inorganic synthetic chemistry is reviewed by focusing on biomimetic synthesis and rationally designed synthesis Focuses on the chemistry of inorganic synthesis assembly and organization of wide ranging inorganic systems Covers all major methodologies of inorganic synthesis Provides state of the art synthetic methods Includes real examples in the organization of complex inorganic functional materials Contains more than 4000 references that are all highly reflective of the latest advancement in inorganic synthetic chemistry Presents a comprehensive coverage of the key issues involved in modern inorganic synthetic chemistry as written by experts in the field

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