

ACS Practice Test 1

Acids & Bases:

1. All are examples of Lewis acid–base reactions *except*

- (A) $\text{Cu}^{2+}(\text{aq}) + 4\text{NH}_3(\text{aq}) \rightleftharpoons [\text{Cu}(\text{NH}_3)_4]^{2+}(\text{aq})$
 (B) $\text{HCl}_{(\text{g})} + \text{NH}_{3(\text{g})} \rightarrow \text{NH}_4\text{Cl}_{(\text{s})}$
 (C) $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightleftharpoons \text{H}_2\text{O}(\text{l})$
 (D) $2\text{Na}_{(\text{s})} + \text{Cl}_{2(\text{g})} \rightarrow 2\text{NaCl}_{(\text{s})}$

2. According to the Lewis definition, an acid is a species

- (A) having a hydrogen ion. (B) donating a pair of electrons.
 (C) accepting a pair of electrons. (D) accepting a hydrogen ion.

3. All are potential Lewis bases *except*

- (A) NH_3 (B) H_2O (C) CH_4 (D) CN^-

4. Which hydrated metal ion is most acidic under conditions of equal molar concentration in water?

- (A) Al^{3+} (B) Ba^{2+} (C) K^+ (D) Zn^{2+} (E) Ag^+

5. Which element in Group 5A forms the most basic oxide?

- (A) N (B) P (C) As (D) Sb (E) Bi

Atomic Structure:

6. In what respect does an atom of magnesium differ from a magnesium ion (Mg^{2+})?

- (A) The ion has a more stable electronic arrangement than the atom.
 (B) The positive charge on the nucleus of the ion is two units greater than the nuclear charge on the atom.
 (C) The ion has two more protons than the atom.
 (D) The ion will react more readily with free Cl_2 .
 (E) The ion has two more planetary electrons than the atom.

7. A certain atom has the symbol $^{148}_{73}\text{Px}$. What does this symbol tell about an atom of the element?

- (A) It has 148 electrons. (B) It has a valence of 4.
 (C) It has 221 protons in the nucleus. (D) It has 75 neutrons in the nucleus.
 (E) Its atomic mass is 73 u.

8. The atomic mass of an element is 32.07 u and its atomic number is 16. The number of protons contained in the nucleus of an atom of this element is

- (A) 16 (B) 31 (C) 32 (D) 48

9. The atomic mass of an element is 32.07 u and its atomic number is 16. The element forms a simple ion. The ion will most likely have a charge of

- (A) 1– (B) 2– (C) 3+ (D) 1+

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J Ma



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Table of Contents Acs Hs13 Exam Answer

1. Understanding the eBook Acs Hs13 Exam Answer
 - The Rise of Digital Reading Acs Hs13 Exam Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Acs Hs13 Exam Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acs Hs13 Exam Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acs Hs13 Exam Answer
 - Personalized Recommendations
 - Acs Hs13 Exam Answer User Reviews and Ratings
 - Acs Hs13 Exam Answer and Bestseller Lists
5. Accessing Acs Hs13 Exam Answer Free and Paid eBooks
 - Acs Hs13 Exam Answer Public Domain eBooks
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 - Acs Hs13 Exam Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acs Hs13 Exam Answer
 - Highlighting and Note-Taking Acs Hs13 Exam Answer
 - Interactive Elements Acs Hs13 Exam Answer
8. Staying Engaged with Acs Hs13 Exam Answer

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Acs Hs13 Exam Answer
- 9. Balancing eBooks and Physical Books Acs Hs13 Exam Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acs Hs13 Exam Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acs Hs13 Exam Answer
 - Setting Reading Goals Acs Hs13 Exam Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acs Hs13 Exam Answer
 - Fact-Checking eBook Content of Acs Hs13 Exam Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
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 - Interactive and Gamified eBooks

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