ACS Practice Test 1

Acids & Bases:

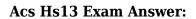
- All are examples of Lewis acid-base reactions except
 - (A) $Cu^{2*}(ng) + 4NH_3(ng) \ge [Cu(NH_3)_4]^{2*}(ng)$
 - (B) $HCl_{(g)} + NH_{3(g)} \rightarrow NH_4Cl_{(i)}$
 - (C) $H^*(og) + OH^*(og) \neq H_2O(i)$
 - (D) $2Nacci + Cl_{2(2)} \rightarrow 2NaClcci$
- 2. According to the Lewis definition, an acid is a species
 - (A) having a hydrogen ion. (B) donating a pair of electrons.
 - (C) accepting a pair of electrons. (D) accepting a hydrogen ion.
- 3. All are potential Lewis bases except
 - (A) NH₃ (B) H₂O (C) CH₄ (D) CN-
- 4. Which hydrated metal ion is most acidic under conditions of equal molar concentration in water?
 - (A) Al3+ (B) Ba2+ (C) K+ (D) Zn2+ (E) Ag+
- 5. Which element in Group 5A forms the most basic oxide?
 - (A) N (B) P (C) As (D) Sb (E) Bi

Atomic Structure:

- In what respect does an atom of magnesium differ from a magnesium ion (Mg²⁺)?
- (A)The ion has a more stable electronic arrangement than the atom.
- (B)The positive charge on the nucleus of the ion is two units greater than the nuclear charge on the atom.
- (C)The ion has two more protons than the atom.
- (D)The ion will react more readily with free Cl2.
- (E)The ion has two more planetary electrons than the atom.
- 7. A certain atom has the symbol ¹⁴⁸₇₃Px. What does this symbol tell about an atom of the element?
 - (A) It has 148 electrons. (B) It has a valence of 4.
 - (C) It has 221 protons in the nucleus.
 (D) It has 75 neutrons in the nucleus.
 - (E) Its atomic mass is 73 u.
- The atomic mass of an element is 32.07 u and its atomic number is 16. The number of protons contained in the nucleus of an atom of this element is
 - (A) 16 (B) 31 (C) 32 (D) 48
- The atomic mass of an element is 32.07 u and its atomic number is 16. The element forms a simple ion. The ion will most likely have a charge of
 - (A) I- (B) 2- (C) 3+ (D) I+

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J Ma



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