

1. According to the Bronsted—Lowry definition, which chemical species can function *both* as an acid and as a base?

17  
 a.  $\text{Cl}^-$   
 b.  $\text{SO}_4^{2-}$   
 c.  $\text{NH}_4^+$   
 d.  $\text{HCO}_3^-$   
 e.  $\text{H}_3\text{O}^+$

2. In the reaction  $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$  which is an acid—base conjugate pair?

17  
 a.  $\text{H}_2\text{O}$  and  $\text{HCN}$   
 b.  $\text{H}_2\text{O}$  and  $\text{OH}^-$   
 c.  $\text{CN}^-$  and  $\text{H}_2\text{O}$   
 d.  $\text{HCN}$  and  $\text{OH}^-$

3. Given that  $\text{HX}$  is a stronger Bronsted acid than  $\text{HY}$  in aqueous solution, which is true of a 1 M solution of  $\text{NaX}$ ?

17  
 a. It is less basic than a 1 M solution of  $\text{NaY}$ .  
 b. It is more basic than a 1 M solution of  $\text{NaY}$ .  
 c. It yields a neutral solution.  
 d. It is more concentrated than a 1 M solution of  $\text{NaY}$ .

4. The pH of a 0.03 M  $\text{HCl}$  solution is

17  
 a. 1.5  
 b. 2.5  
 c. 3.5  
 d. 12.5

5. The pH of a  $1.0 \times 10^{-3}$  M  $\text{Ba}(\text{OH})_2$  solution at  $25^\circ\text{C}$  is

17  
 a. 2.7  
 b. 3.0  
 c. 11.0  
 d. 11.3

6. Which is the strongest acid?

17  
 a.  $\text{HClO}$   
 b.  $\text{HClO}_2$   
 c.  $\text{HClO}_3$   
 d.  $\text{HClO}_4$

7. Which salt reacts with water (hydrolyzes) to produce a basic solution?

17  
 a.  $\text{NaC}_2\text{H}_3\text{O}_2$   
 b.  $\text{NaNO}_3$   
 c.  $\text{NH}_4\text{Cl}$   
 d.  $\text{BaSO}_4$

8. The *weakest* of the bases listed is

Acid	Conjugate Base	$K_a$ (Ionization Constant of Acid)
$\text{HCl}$	$\text{Cl}^-$	100% ionized
$\text{HSO}_4^-$	$\text{SO}_4^{2-}$	$1.2 \times 10^{-2}$
$\text{H}_2\text{S}$	$\text{HS}^-$	$5.7 \times 10^{-8}$
$\text{HS}^-$	$\text{S}^{2-}$	$1.2 \times 10^{-14}$

17  
 a.  $\text{Cl}^-$   
 b.  $\text{CN}^-$   
 c.  $\text{HS}^-$   
 d.  $\text{S}^{2-}$   
 e.  $\text{SO}_4^{2-}$

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Periodic Table									
1	2	3	4	5	6	7	8	9	10
11	12	13	14	15	16	17	18	19	20
21	22	23	24	25	26	27	28	29	30
31	32	33	34	35	36	37	38	39	40
41	42	43	44	45	46	47	48	49	50
51	52	53	54	55	56	57	58	59	60
61	62	63	64	65	66	67	68	69	70
71	72	73	74	75	76	77	78	79	80
81	82	83	84	85	86	87	88	89	90

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10. Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate  $\text{KC}_2\text{H}_3\text{O}_2$  is less alkaline than a 0.10 M solution of potassium cyanide,  $\text{KCN}$ ?

17  
 a. Hydrocyanic acid is a weaker acid than acetic acid.  
 b. Hydrocyanic acid is less soluble in water than acetic acid.  
 c. Cyanides are less soluble than acetates.  
 d. Acetic acid is a weaker acid than hydrocyanic acid.

11. In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M  $\text{NaOH}$ , the properties of the solution at the equivalence point will correspond exactly to the properties of

18  
 a. a 0.100 M sodium solution.  
 b. a 0.0500 M sodium hydroxide solution.  
 c. a 0.0500 M benzoic acid solution.  
 d. a 0.0500 M sodium benzoate solution.

# Acs Gob Chemistry Practice Exam

**William W. Aljoe, Robert R. Stein, R. C.  
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