

ABBREVI	ATTION	S AND SYMBOLS		
amount of substance		joule	3.	
atmosphere	atem	kelvin	180	
boiling pt. elev. const.	Kee	liter	, E.	
Celsius temperature	~C	molal	199	
energy of activation	E_{α}	molar	0.40	
enthalpy	PF	mole	molt	
entropy	5	pressure	P	
equilibrium constant	King	second	56	
free energy	6	specific heat capacity	Co	
freezing pt. depr. const.	Kir.	temperature	T	
gram	- 65	wolt	14°	
hour	The second	volume	W	

CONSTANTS								
Planck's constant h = 6.626×10 ⁻³⁴ J-s	Gas Constant R = 0.0821 (L-atm)/(mol-K) R = 8.314 J/(mol-K)							
Speed of light c = 2.998×10 ⁸ m·s ⁻¹ Rydberg constant	Water Dissociation Constant K _w = 1.0×10 ⁻²⁴ at 25 °C							
R _H = 2.18×10 ⁻¹⁸ J Avocadro Constant	Faraday Constant F = 96470 C/mol e ⁻							
$N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$	F = 96470 A-s/mol e ⁻ F = 96470 J/V-mol e ⁻							
Measures of pressure 1 atm = 760 mmHg								

	EQ	UATIONS			
$E = hv$ $E = \frac{hc}{\lambda}$	$\frac{\text{rote}_1}{\text{rote}_2} = \sqrt{\frac{\text{molar mass}_2}{\text{molar mass}_1}}$ $\Delta T_1 = K_1 m \qquad \Delta T_2 = I K_1 m$	Integrated Rate Laws. zero: $[A] = [A]_c - \hat{x}t$	$pH = pK_x + log \left(\frac{A}{[HA]} \right)$ $\Delta G^{\circ} = \Delta H^{\circ} - T\Delta S^{\circ}$		
$\Delta \mathcal{E} = R_n \left[\frac{1}{\alpha_i^2} - \frac{1}{\alpha_f^2} \right]$ $q = m \in \Delta T$ PV = nRT	$\Delta T_b = K_b m$ $\Delta T_b = I K_b m$ $\Omega = AART$ $\Omega = I AART$ $ln\left(\frac{k_2}{k_1}\right) = \frac{\mathcal{E}_d}{R} \left(\frac{1}{T_1} - \frac{1}{T_2}\right)$	second: $\frac{1}{[A]} = \frac{1}{[A]_0} + kt$ first: $\ln[A] = \ln[A]_0 - kt$	$\Delta G^{o} = -RT \ln K$ $E^{o} = \frac{RT}{\alpha E} \ln K$	$\Delta G = \Delta G^{\circ} + RT \ln \epsilon$ $\mathcal{E} = \mathcal{E}^{\circ} - \frac{RT}{n^{p}} \ln Q$	

1				P	ERIC	DIC	TAE	BLE (OF T	HE	ELEN	MEN"	TS				18
1 H 1.000	2											13	14	15	16	17	2 He 4.000
3 Ui	4 Be											5 B 10.81	6 C	7 M 14 01	8 O 16.00	9 F 23.00	10 Ne
11 No 22.99	12 Mg 24.30	3	4	5	6	7	8	9	10	11	12	13 Al 20.00	1.4 Si 28.09	15 P 10.67	16 \$ \$2.07	17 Cl 35.45	18 Ar 19.55
19 K 96.10	20 Ca ac.os	21 Sc 44,96	22 11 47.88	23 V 50.94	24 Cr 53.00	25 Mn 54.94	26 Fe 55.45	27 Co 58.93	28 Ni 18.69	29 Cu 63-53	30 Zn 65.39	31 Ga 49.72	32 Ge 72.61	33 As 74/9	34 Se 78.96	35 Br 29.90	36 Kr 81.80
37 Rb	38 \$r 87.63	39 Y	40 Zr 91,33	41 Nb	42 Mo	43 Tc	44 Ru 101.1	45 Rh 101.9	46 Pd 105.4	Ag Lora	48 Cd 113.4	49 In	50 Sn 118.7	51 Sb 131.0	52 Te	53 1 126.9	5-4 Xe 131.3
55 Cs 130.0	56 Ba	57 La	72 Hif 176.5	73 Ta	74 W	75 Re 186.2	76 Os 1903	77 Ir 193.3	78 Pt 195.1	79 Au 197.0	80 Hg 200.6	8:1 T1 204.4	82 Pb 207.2	83 84 209.0	84 Po	85 At	86 Rn
87 Fr	88 Ra	89 Ac	104 Rf	105 Db	106 Sg	107 Bh	108 Hs	109 Mt	110 Ds	111 Rg	112 Cn	113 Nb	114 FI	115 Mc	116 Lv	117 Ts	118 Og
		58 Ce	59 Pr	60 Md	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	71 Lu	1	
		90 Th	91 Pa 233.0	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr]	

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 $buffer\ ph\ 2\ water\ bath\ pipettes\ glass\ rods\ sharpers\ container\ wash\ bin$

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web action of a digestive enzyme purpose of the exercise to investigate the action of amylase and the effect of heat on its enzymatic activity materials needed 0 5 amylase solution beakers 50 and 500 ml distilled water funnel pipettes 1 and 10 ml pipette rubber bulbs 0 5 starch solution graduated cylinder

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