

Acids & Bases Review

1. Name the following compounds as acids and circle which are weak acids:



nitric acid



sulfuric acid



hydrofluoric acid



carbonic acid



acetic acid



phosphoric acid

2. Write the formulas for the following bases:

a) calcium hydroxide



b) aluminum hydroxide



c) ammonium hydroxide



d) lithium hydroxide



3. Complete the following table:

	Acids	Bases
Definition	A substance that increases the concentration of H^+ (or H_3O^+) when added to water. or A substance that acts as a proton (H^+) donor.	A substance that increases the concentration of OH^- when added to water. or A substance that acts as a proton (H^+) acceptor.
Electrolyte in aqueous solution?	H^+	OH^-
pH	less than 7	greater than 7
Taste	sour	bitter
Litmus turns	blue turns red	red turns blue
Which reacts with metals to form H_2 gas?	yes	no

Conjugate acid-base pairs:

4. Identify the acid, the base, the conjugate acid and the conjugate base in the reactions:



5. An acid will form the conjugate base and a base will form the conjugate acid.

Acids Bases And Solutions Chapter Test Answers

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Table of Contents Acids Bases And Solutions Chapter Test Answers

1. Understanding the eBook Acids Bases And Solutions Chapter Test Answers
 - The Rise of Digital Reading Acids Bases And Solutions Chapter Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Acids Bases And Solutions Chapter Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acids Bases And Solutions Chapter Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acids Bases And Solutions Chapter Test Answers

- Personalized Recommendations
 - Acids Bases And Solutions Chapter Test Answers User Reviews and Ratings
 - Acids Bases And Solutions Chapter Test Answers and Bestseller Lists
5. Accessing Acids Bases And Solutions Chapter Test Answers Free and Paid eBooks
 - Acids Bases And Solutions Chapter Test Answers Public Domain eBooks
 - Acids Bases And Solutions Chapter Test Answers eBook Subscription Services
 - Acids Bases And Solutions Chapter Test Answers Budget-Friendly Options
 6. Navigating Acids Bases And Solutions Chapter Test Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Acids Bases And Solutions Chapter Test Answers Compatibility with Devices
 - Acids Bases And Solutions Chapter Test Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acids Bases And Solutions Chapter Test Answers
 - Highlighting and Note-Taking Acids Bases And Solutions Chapter Test Answers
 - Interactive Elements Acids Bases And Solutions Chapter Test Answers
 8. Staying Engaged with Acids Bases And Solutions Chapter Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acids Bases And Solutions Chapter Test Answers
 9. Balancing eBooks and Physical Books Acids Bases And Solutions Chapter Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acids Bases And Solutions Chapter Test Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Acids Bases And Solutions Chapter Test Answers
 - Setting Reading Goals Acids Bases And Solutions Chapter Test Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Acids Bases And Solutions Chapter Test Answers

- Fact-Checking eBook Content of Acids Bases And Solutions Chapter Test Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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