Quiz 2: Acids and Bases Review Sheet

Acids and Bases:

- Many scientists tried to define what is an acid see the table below
- Amphoteric substance is a substance that behaves both as an acid or a base; examples: H2O, HCO3/HSO4/HPO4⁻².....
- An acid can be described as a one that gives a proton H+ or a hydronium ion HsO* in an
 aqueous solution.
- Acids can be classified according to the maximum number of protons given in a solution:
- Monoprotic: one proton as CH3COOH, HCl, HCN
- Diprotic: two protons: H₂SO₄, H₂CO₃
- Triprotic: three protons H3PO4
- Salt formation: When an acid and a base react; it is a double replacement reaction and the
 products are salt and water.

Acids	Bases	
Taste is sour	Taste is bitter	
Pungent and corrosive	Slippery feeling and corrosive	
Change litmus paper to red	Change litmus paper to blue	
pH <7	pH >7	
[H ₂ O ⁺] > [OH ⁻]	[H ₂ O*] < [OH·]	

	Acid	Base	Notes
Arrhennus	Hydrogen ion producer	OH Producer	Depend on structure should have H:O*/OH in its formula
Brosted- Lowery	Proton/hydrogen ion donor	Proton/hydrogen ion acceptor	Its acids are the same as Arrhenius Bases include NH3 (ammonia) and organic bases that couldn't fir Arrhenius

Acids And Bases Review Answer Sheet

WJ Hussar

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