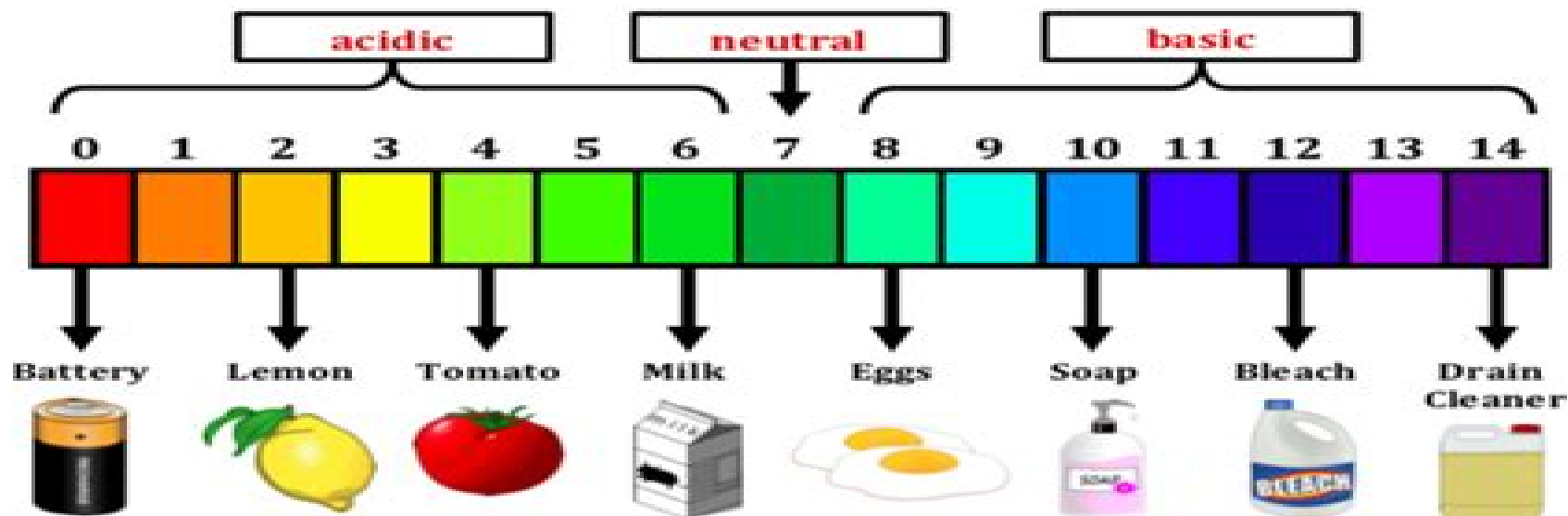


# Introduction to the pH Scale

The pH scale is a measure of the **hydrogen ion** ( $H^+$ ) concentration of a substance. It is used to indicate the **acidity** of a solution.

- **Acids** have a pH **below 7**. The lower the pH, the more acidic.
- **Bases** have a pH **above 7**. The higher the pH, the more basic.

**Instructions:** In the pH diagram, label the regions as **acidic**, **basic**, or **neutral** ( $H_2O$ ).



**Instructions:** Answer the questions below regarding the pH scale.

- 1) Which is more acidic: **milk** or **lemon**?
- 2) Which is more basic: **soap** or **bleach**?
- 3) What is the only **neutral** substance (not acidic or basic)?
- 4) What is the strongest **acid** shown on the pH scale above?
- 5) What is the strongest **base** shown on the pH scale above?

**lemon**

**bleach**

**water**

**battery**

**drain cleaner**

# Acids And Bases Practice Worksheer

**T. R. Dickson**



## Acids And Bases Practice Worksheer:

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