

Measures of Acidity and Basicity

- pH and pOH
 - $-\log[H^+]$ and $-\log[OH^-]$ respectively
 - **Typically** between 1 and 14
 - 7 is neutral
 - **Smaller pH = more acidic**
 - **Larger pH = more basic** (just remember to 'turn up the base!')
 - $[H^+ \text{ or } OH^-] = e^{-pH \text{ or } pOH}$
- $K_{a \text{ or } b}$
 - The larger the number, the stronger the acid or base (depending on a or b)

Acidity Basicity And Ph Concept Review

Lawrence P. Bevy



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Tarendash, 2021-01-05 Barron's Let's Review Regents Chemistry gives students the step by step review and practice they need to prepare for the Regents Chemistry Physical Setting exam. This updated edition is an ideal companion to high school textbooks and covers all Chemistry topics prescribed by the New York State Board of Regents. Let's Review Regents Chemistry covers all high school level Chemistry topics and includes Extensive review of all topics on the test. Extra practice questions with answers. A detailed introduction to the Regents Chemistry course and exam. One actual recently released Regents Chemistry exam with an answer key. **Publications of the National Bureau of Standards** United States. National Bureau of Standards, 1970 **NBS Special Publication**, 1968 **Publications** United States. National Bureau of Standards, 1969 *Publications of the National Institute of Standards and Technology ... Catalog* National Institute of Standards and Technology (U.S.), 1970 Holt Chemistry Holt Rinehart & Winston, 2003-01-24 *Catalog of National Bureau of Standards Publications, 1966-1976* United States. National Bureau of Standards, 1978 Publications of the National Bureau of Standards, 1968-1969 United States. National Bureau of Standards, Betty L. Oberholtzer, 1970

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