

## Chem 160 Acid-Base Equilibria (key)

1. Write an equilibrium equation for each of the following **acids**. Identify the acid (a), base (b), conjugate acid (ca) and conjugate base (cb).



2. Write an equilibrium equation for each of the following **bases**. Identify the acid, base, conjugate acid and conjugate base.



3. Write the chemical formula for the conjugate base of each acid.



3. Write the chemical formula for the conjugate acid of each base.



4. For the following amphoteric (amphiprotic) compounds, write an equilibrium equation to show how the compound can act as an acid and another equilibrium equation to show how the compound can act as a base.

# Acid Base Equilibria Pogil Answer Key

**Judd E. Hollander**



## **Acid Base Equilibria Pogil Answer Key:**

*Analytical Chemistry* Juliette Lantz, Renée Cole, The POGIL Project, 2014-12-31 An essential guide to inquiry approach instrumental analysis Analytical Chemistry offers an essential guide to inquiry approach instrumental analysis collection The book focuses on more in depth coverage and information about an inquiry approach This authoritative guide reviews the basic principles and techniques Topics covered include method of standard the microscopic view of electrochemistry calculating cell potentials the BerriLambert atomic and molecular absorption processes vibrational modes mass spectra interpretation and much more *Making Chemistry Relevant* Sharmistha Basu-Dutt, 2010-02-19 Unique new approaches for

making chemistry accessible to diverse students Students interest and achievement in academics improve dramatically when they make connections between what they are learning and the potential uses of that knowledge in the workplace and or in the world at large Making Chemistry Relevant presents a unique collection of strategies that have been used successfully in chemistry classrooms to create a learner sensitive environment that enhances academic achievement and social competence of students Rejecting rote memorization the book proposes a cognitive constructivist philosophy that casts the teacher as a facilitator helping students to construct solutions to problems Written by chemistry professors and research groups from a wide variety of colleges and universities the book offers a number of creative ways to make chemistry relevant to the student including Teaching science in the context of major life issues and STEM professions Relating chemistry to current events such as global warming pollution and terrorism Integrating science research into the undergraduate laboratory curriculum Enriching the learning experience for students with a variety of learning styles as well as accommodating the visually challenged students Using media hypermedia games and puzzles in the teaching of chemistry Both novice and experienced faculty alike will find valuable ideas ready to be applied and adapted to enhance the learning experience of all their students

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**Acid Base Equilibria** E. King,R. Robinson,1965-01-01 *Acid-base Equilibria* Edmund James King,1965  
*Acid-base Equilibria* Edward J. King,1965 **Science Citation Index** ,1993 Vols for 1964 have guides and journal lists  
**ACID BASE EQUILIBRIA.** EJ. KING,1965 **Acid-Base Equilibria in Aqueous and Nonaqueous Solutions** Martin Kilpatrick,1947 For uncharged proton acids in basic solvents of not too low dielectric constant the reaction  $Ax S \rightleftharpoons SH Bx$  1 takes place the extent of the reaction depending on the strength of the acid  $Ax$  relative to the basic strength of the solvent  $S$  The equilibrium constant for this reaction  $K_{AxS} = \frac{[SH][Bx]}{[Ax][S]}$  is usually expressed as the more commonly used dissociation constant  $K_c$  which includes the concentration of the solvent  $K_{AxS} = K_c [S]$  3 The constant  $K_c$  varies with ion concentration and in the limited range of concentration where the Debye H ckel theory applies the following equation may be used to determine the thermodynamic equilibrium constant  $K_c \log K_c = \log K_c^o - \frac{A}{2} \sqrt{I}$  4 where  $A = \frac{2.303}{DkT} \left( \frac{1}{\epsilon} - \frac{1}{\epsilon_0} \right)$  2  $\frac{1}{\epsilon} = \frac{1}{\epsilon_0} + \frac{4\pi N}{9kT} \left( \frac{e^2}{\epsilon_0} \right)$  2 The values for  $A$  at 25 C are for water dielectric constant  $D = 78.54$  1 020 for methyl alcohol  $D = 31.54$  02 for ethyl alcohol  $D = 24.25$  97 and for butyl alcohol  $D = 17.49$  79 For more concentrated solutions no theoretical equation can be given but Table I shows that the change in the equilibrium constant is of considerable magnitude [Acid-base Equilibria in Nonaqueous Solvents](#) Joseph Anthony Caruso,1967 **Aqueous Acid-base Equilibria and Titrations** Robert De Levie,2023 This text will give students a thorough grounding in P H and associated equilibrium material absolutely fundamental to the understanding of many aspects of chemistry This book uses the new theoretical developments that have led to more generalized approaches to equilibrium problems which are often simpler than the approximations which they replace *Acid-base equilibria in nonpolar media* Eva-Ingrid Rööm,2007 **Acid-base Equilibria in Anhydrous Formic Acid** John Clifford Marshall,1960 *Acid-base Equilibria in Non-aqueous Solvents* ,1958 **Acid-base Equilibria in Concentrated Salt Solutions** Augustus Donovan Saxton,1968 [Acids and Bases](#) Ronald Percy Bell,1952 **Acid-base Equilibria, with Special Reference to Indicators for Use in Non-aqueous Media. Indicators for Non-aqueous Titration of Weak Acids** A. Tetlow,1965

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