

# Solutions, acids, and bases

## 1. VOCABULARY REVIEW OF SOLUTIONS

**Solute** - what is **dissolved** in the solvent to form a solution.

**Solvent** - the substance that **dissolves** the solute, ordinarily a liquid when the solute and the solvent come together to form a **homogeneous mixture**.

**Miscible** - two mix in any proportions (none to be mixed) - liquids will pour each over

**Immiscible** - don't mix in any proportions (none to be mixed)

**Examples of miscible-immiscible** - ethanol-water, oil-water, oil-water, oil-water

**Solute** - **amount** of **being dissolved** in liquid

**Insoluble** - a substance (tends to be a solid) that **will not dissolve** in a solvent (an example is sand and water)

**Suspension** - looks like a **solution** when **stirred**, the **solutes** a bit of water (suspensions) also it is a mixture - gaseous mixture where the small solid particles of a substance are spread throughout a liquid without settling.

**Unsaturated** - **solids** with **solubility chart** when the temperature would be enough to make a substance dissolve but at a greater amount of **solute** - **solubility less** - **solute** **exceeds** what can be dissolved (e.g. salt & water)

**Saturated** - **solids** with **solubility chart** when the temperature-solubility relationship is at the line in other words a **solution** with **solute** has **reached** or the **maximum possible amount**.

**Supersaturated** - contains **more than the max. amount** of **solute** that can be dissolved at the temperature on the solubility chart this would mean the grams of solute would go above the line.

## 2. SOLUBILITY VS. TEMPERATURE CHART



Above the line - **Supersaturated**

At the line - **Saturated**

Under the line - **Unsaturated**

## 3. METHODS OF SEPARATION

**Filtration** - separate particles by **size**

**Distillation** - separate by **boiling point** - **more** and **liquids** to be separated.

**Extraction** - allows one to remove a **liquid from a solid** by **transferring** the **liquid** to a **gaseous phase**.

**Centrifugation** - separating molecules by their **densities** by **spinning** them **quick**.

**Decanting** - process of separating **insoluble** or **suspension** solutions by **pouring** away the **liquid** from the **solid**.

**Chromatography** - separation by the **rates** at which components of different solutions **move through** a **medium** (e.g. gel electrophoresis)

# Acid And Bases Study Guide

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