

ACID-BASE TITRATION

Name _____

An Acid-Base Titration is the determination of the concentration of an acid or base by exactly neutralizing the acid or base with an acid or base of known concentration.

Example: It takes 45.2 ml. of 0.1 M hydrochloric acid to neutralize 30 ml. of sodium hydroxide. What was the concentration of sodium hydroxide?



Step 1: Express the known concentration in liters .045 L of HCl

Step 2: Multiply by its concentration: (.045 L)(0.1 M) = .0045

Step 3: Multiply by the mole to mole comparison (see equation)
(.0045) (1/1) = .0045 still

Step 4: Divide by the unknown concentration in liters

$$.0045 / .030 \text{ L} = .150 \text{ M}$$

1. In the titration of 10 ml of vinegar 40 ml of .1 M sodium hydroxide was used. What was the molarity of the vinegar?



2. If 40 ml. of .1 M sodium hydroxide neutralizes 30 ml. of HCl, what was the molarity of the HCl?



3. It took 18 ml. of .2 molarity to titrate 45 ml. of $\text{Ca}(\text{OH})_2$. What was the molarity of the $\text{Ca}(\text{OH})_2$?



4. 60 ml. of .5 M of HCl is needed to neutralize 60 ml. of NaOH. Calculate the molarity of the NaOH



5. To titrate 30 ml. of potassium hydroxide required 20 ml. of .1 M Hydrochloric Acid. Write the equation and tell the molarity of the potassium hydroxide.

Acid Base Titration Worksheet With Answers

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