

## Acids and bases review sheet Answer key

1. List the differences between an acid and a base. Your description should include such things as whether both are electrolytes, which contains a hydrogen ion and which includes a hydroxide ion, and you should also include two examples of an acid and two examples of a base that are present in your household.

Acid	base
Conducts electricity Contains hydrogen ion Proton donor Turns litmus paper red React with most metals Releases hydronium Ex: lemon Vinegar Battery acid (sulfuric acid)	conducts electricity contains hydroxide ion proton acceptor turns litmus paper blue releases hydroxide ion Ex: Windex bleach soap

2. Listed below are some of the properties of acids and bases. Fill in the blanks with the appropriate word, **acids** or **bases**:

- A**   produce hydrogen ions ( $H^+$ ) in solution
- B**   have a bitter taste
- B**   produce hydroxide ions ( $OH^-$ ) ions
- A**   react with bases to form salts
- A**   have a sour taste
- B**   have a slippery, soapy feel
- B**   react with acids to form salts
- A**   react with many metals

3. Write an equation for the dissociation of the following **acids** in water: follow example

- |   |                                     |
|---|-------------------------------------|
| a) $HClO_4 \rightarrow H^+ + ClO_4^-$         | d) $H_2S \rightarrow 2H^+ + S^{2-}$ |
| b) $H_2SO_4 \rightarrow 2H^+ + SO_4^{2-}$     | e) $HCl \rightarrow H^+ + Cl^-$     |
| c) $HC_2H_3O_2 \rightarrow H^+ + C_2H_3O_2^-$ | f) $HNO_3 \rightarrow H^+ + NO_3^-$ |

4. Write an equation for the dissociation (ionization) of the following **bases** in water: (follow example)

- |   |   |
|---|---|
| a) $Mg(OH)_2 \rightarrow Mg^{2+} + 2OH^-$   | d) $KOH \rightarrow K^+ + OH^-$             |
| b) $NaOH \rightarrow Na^+ + OH^-$           | e) $Mg(OH)_2 \rightarrow Mg^{2+} + 2(OH^-)$ |
| c) $Ba(OH)_2 \rightarrow Ba^{2+} + 2(OH^-)$ | f) $LiOH \rightarrow Li^+ + OH^-$           |

5) A. What type of reaction is a **neutralization** reaction? **Double replacement**

B. Write the net equation for neutralization:



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**The Princeton Review**



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