- Name the following compounds as acids and circle which are weak acids:
- a) HNOs b) H<sub>2</sub>SO<sub>4</sub> c) HF d) H<sub>2</sub>CO<sub>3</sub> e) HC<sub>2</sub>H<sub>1</sub>O<sub>2</sub>
- f) H<sub>1</sub>PO<sub>4</sub>
- 2. Write the formulas for the following bases a) calcium hydroxide b) aluminum hydroxide c) ammonium hydroxide d) lithium hydroxide

3. Complete the following table:

	Acids	Bases
Defination	3333333	100.000
Electrolyte in equeous solution?		
pH		
pH Taste		
Laterous towns		
Which reacts with metals to form H; gas?		

Conjugate acid-base pairs

- 4. Identify the acid, the base, the conjugate acid and the conjugate base in the reaction  $NH_1(g) + HOH(l) \Rightarrow NH_1(an) + OH(an)$
- Identify the acid, the base, the conjugate acid and the conjugate base in the reaction: HCO<sub>1</sub> (an) + HOH (l) ≠ CO<sub>1</sub> (an) + H<sub>1</sub>O\* (an)
- 6 An acid will form the conjugate and a base will form the conjugate

Neutralization and Titration.

- 7. Classify the following acids as monoprotic, diprotic or triprotic a) H<sub>2</sub>SO<sub>4</sub> b) H<sub>2</sub>PO<sub>4</sub> e) HCI d) HC1H1O2 or CH1COOH (same thing)
- 8 What is the concentration of an unknown NaOH solution if took an average of 15 4mL of 0.100N HCl solution to neutralize 10.0mL of the NaOH solution?

From the given data, what is the molarity and the normality of the HF solution if the KOH solution. WAR D SDM7

	Acid, HF	Base, KOH
Initial Volume, mL	5.5	14.9
Final Volume, mL	25.9	24.4

pH = - log[H<sub>2</sub>O'] To find pH if given [H<sub>2</sub>O'], enter the [H<sub>2</sub>O'], press Log & then the '+ button To find [H:O'] if given pH, enter negative pH, then 2nd function Log

10 A solution has a [H<sub>2</sub>O\*] of 1 X 10.7 M, what is its pH? Is this an acid, base or is it neutral?

TURN THE PAPER OVER

11. What is the pH of a solution if the [H\*] = 7.2 X 10°M? Is this an acid or base?

# **Acid And Base Answer Key**

**Disha Experts** 

#### **Acid And Base Answer Key:**

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