# General Chemistry Reference Sheet

This reference sheet addresses some of the more psychiar pieces of information that need to be memorized in a general chemistry roome. It also contains a simple set of exsential formedas in chemistry with continue, explanations, and general tips.

This sheet is pursuit to be as concine as possible, and many information in the textbook is left out in favor of runtions and tips. This short is, therefore, best used as a supplemost to, ant a replacement of the textbook.

#### SI Fundamental Units

Mass	Kologram (kg)
Length	Motivi (m.)
Timer	Societal (s)
Temperature	Kelvin (K)
Amount of substance	Mole (med)
Electric current	Ampere (A)
Londoon intensity	Candida (nd)

#### Atomic Experiments and Models

J. J. Thomson Discovered e"; Cathode ver Plean probling model B. A. Millian Measured charge of e ... Oil deep H. Berganet/M. Curie Discovered radioactivity E. Butherland. Discovered is, 3, and 5 more Discovered nucleon: Gold foil experiment J. Chadwick Discovered neutrons. N. Budar Bloke model (electron orbits) Chaptering mechanicists: Ouantum model:

#### Polyatomic Ions

C. Marine Street, Street, Street, St.	DEC. RESIDEN		
2014.	Antibonism	ORC	hydroxide
CNT	cyantide.	$C_2O_2^{p-}$	occulator
$O_2^{p-}$	peruside	CNO	vysoume:
BESON"	hydrogen ind-	C <sub>2</sub> H <sub>2</sub> O <sub>2</sub> **	postate.
	fate		
SKINT	thiocymate	NO.	sultinate.
80,0	sulfite	C10,	perchlorate
$CO_{h^{2\alpha}}$	carbonate	C1O <sub>2</sub> "	chlorate
PO4**	phosphate	C107	Arporblorite
$S_2O_1^{-2-}$	thinedlete	HIPO, T	bordrogen
			phosphate
CrO <sub>2</sub> 3-	(biromake	H <sub>2</sub> O <sup>4</sup>	Applications
ChyOy3-	dicheoniste	TO <sub>2</sub>	phosphite
MaO <sub>y</sub>	permanganate	Har	mentagy (II)
20.5	workle	Cylin	curbide.
$C_nH_nO_n^{-1}$	factories -	80	diesthicke
O <sub>x</sub>	superoudle	AsO <sub>2</sub> <sup>3</sup>	anomalite.
PO <sub>2</sub> <sup>k</sup>	hypophosphite	AldOs 11	agramada:
SiO <sub>2</sub> 2-	silicator	PyOy**	pyrophoplat

#### Ionic Solubility Chart.

Subable	Exceptions
2005	None
CIE/COO"	None
CI	Ag*, Hg <sub>2</sub> **, P63**
Be	Ag*, Hgp**, P6**
1"	Ac*, Hee**, Ph**
50,1	Se <sup>3+</sup> , Ba <sup>3+</sup> , B <sub>23</sub> <sup>3+</sup> , Pa <sup>2+</sup>
Trackette	Europtions
677	NH <sub>4</sub> °, alkali metal cutions, Ca <sup>21</sup> , Sr <sup>21</sup> , Ba <sup>22</sup>
CO <sub>3</sub> 3	NH,", alkali metal cutions
POv <sup>2</sup>	NH, *, alkali metal cutions
OH!"	NH,", alkali metal cutions, Ca <sup>2+</sup> ,
	Sr <sup>2+</sup> , Ba <sup>2+</sup>

#### Strong Acids and Bases

Strong scids and bases dissociate in water completely.

Strong Arids		Strong Bases
HCI HBe HI HCNO <sub>2</sub>	HCSO <sub>4</sub> HCSO <sub>4</sub> H <sub>2</sub> SO <sub>4</sub> Ra(OH) <sub>2</sub>	Allali metal hydroxides Ca(OH) <sub>2</sub> Sc(OH) <sub>2</sub>

#### Activity Series

Metals below III cannot react with with acids to form III. More active metals are better reducing agents.

#### From most active to least active:

Li\*, K\*, Ba<sup>3+</sup>, Ca<sup>3+</sup>, Na\*, Mg<sup>2+</sup>, Al<sup>3+</sup>, Ma<sup>3+</sup>, Za<sup>3+</sup>, Cr<sup>3+</sup>, Fr<sup>2+</sup>, Co<sup>3+</sup>, Ni<sup>3+</sup>, Sa<sup>2+</sup>, Ph<sup>3+</sup>, H\*, Ca<sup>2+</sup>, Ag\*, Hg<sup>3+</sup>, Pr<sup>3+</sup>, Au<sup>3+</sup>

#### Flame Colors

Calctions	Deick red
Copper (I)	Blue
Copper (II)	Groves or blue-green.
Potavioris	Liber
Lithinse	Dark red.
Nodimm	Hright pellow
Strongtonia	Bird
Bartoni	Light given
Jeon (IIII)	Gold
Cestam	Blue Violet
Deschrosse	Blue
Lead	Blue
Rubidises	Roel-Violet

#### Phase Changes

	From rolld	From liquid	From gar
To sold		Treeslang.	deposition.
To liquid	melting	A STATE OF THE STA	condensation
The same	ent-limetion.	State of Southern	400

#### Section Colored

Dise
George
Purple
Yellow
Onwage
Light liber
aty jelior

#### Thermodynamic Laws

First Law: Energy cannot be created nor destroyed. It can only be transferred in the form of either heat or work. Second Law: Any spontaneous reaction increases the entrease of the universe.

Third Law: An ideal solid crystal at 6 K has an entropy

#### Thermodynamic Formulas

Standard thermodynamic condition	is 208 K; Latin; LM.
Kinetic energy	$K = me^2/2$
Electrostatic potential energy	$E_{R} = (4c_{1}Q_{1}Q_{2})/d$
Internal energy	$\Delta E = q + w$
Eartholps	M = E + FV
Specific beat	$s = q/(m \cdot \Delta T)$
Entropy in revenible reaction.	$\Delta S_{\text{system}} = (\Delta N)/T$
$\Delta \hat{\kappa}$	$m_{\rm constant} = -(\Delta N)/T$
Microstate-entropy relationship	$S = E \ln W$
Gibbs fore energy	G = H - TS
Gibbs free energy change	$\Delta G = \Delta H - T\Delta S$
	$\Delta G = 2\epsilon G^* + RT \ln Q$
March Committee	No. 201

#### Constants

The second secon	
Boltzmann's constant. Ly-	$= 1.381 \times 10^{-23} \text{ m}^3 \text{kg} \cdot \text{s}^{-1} \text{K}^{-1}$
Coulomb's constant by =	$1/(4\pi t_0) = 8.988 \times 10^6 \text{ J} \cdot \text{m/C}^4$
Assignative's assurber	$N_A = 6.022 \times 10^{23} \text{ neol}^{-1}$
Fernelog's executant	$E = 9.649 \times 99^4$ C/mod
Planck's constant	$A = 6.026 \times 10^{-34} \text{ J}_{\odot}$
Ideal gas constants	H = 0.0821  (L.: atm)/(molK)
	R = 8.334 J/(noil : K)
Vacuum permittivity x <sub>0</sub> =	$1/(m_0c^2) \approx 8.854 \times 10^{-11} \text{ F/m}$
Vacuum permeability	$p_0 = 1.297 \times 10^{-6} \text{ N/A}^{-6}$
Atomic mass	$1.8800 = 1.061 \times 10^{-16}$ g
Electron charge	$a = 1.002 \times 10^{-10} \text{ C}$
Electropyolt	$1 \text{ eV} = 1.002 \times 10^{-18} \text{ J}$
Atmospheric pressure	3 atm = 1.013 × 10 <sup>5</sup> Pa
Absolute zero	0 K > 473.15 °C
Spend of light in vacuum	$r \sim 2.598 \times 10^9 \text{ m/s}$

#### Quantum Mechanical Formulas

Energy of a quantum		E = tor
Wavelength-frequency rela-		$x = y \cdot \lambda$
Probability distribution	Pv = []]	$ \psi(x,y,z) ^2 dxdydz$

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**Robert Balay, Eugene Paul Sheehy** 

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