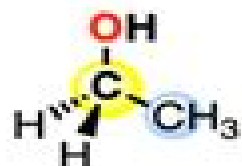
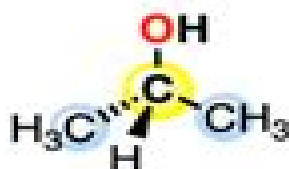


Alcohols are classified as primary (1°), secondary (2°), or tertiary (3°) according to the number of carbons **directly attached** to the carbon bonded to the OH.



Primary (1°) alcohol

- C-OH carbon attached to **one** carbon



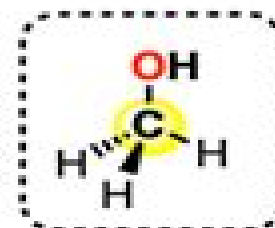
Secondary (2°) alcohol

- C-OH carbon attached to **two** carbons



Tertiary (3°) alcohol

- C-OH carbon attached to **three** carbons



Methanol (unique)

- C-OH carbon attached to **zero** carbons

- The O-H bond is strongly **polarized** due to the large difference in electronegativity between oxygen (3.4) and hydrogen (1.0).



- One consequence of this polarization is **hydrogen bonding**, which is responsible for **strong intermolecular interactions** in alcohols, resulting in high **boiling points** (*relative to alkanes and ethers*)



1-propanol
MW 60.1
bp 97°C



Ethylene glycol
MW 62
BP 197°C

contrast
with



methoxyethane
MW 60.1
bp 8°C

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WJ Hussar



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