

# General Chemistry cheat sheet

## DAT General Chemistry Equation Sheet

### Chapter 0: General and Lab Concepts Review

|                                           |                                           |                                                                                                                     |
|-------------------------------------------|-------------------------------------------|---------------------------------------------------------------------------------------------------------------------|
| <b>Dilutions</b>                          | $M_1V_1 = M_2V_2$ or<br>$C_1V_1 = C_2V_2$ | $M$ or $C$ = concentration<br>$V$ = volume                                                                          |
| <b>Percent Error</b>                      | $\frac{(A - T)}{T} \times 100$            | $T$ = theoretical<br>$A$ = actual                                                                                   |
| <b>Absorbance<br/>(Spectrophotometer)</b> | $Abs = \epsilon cl$                       | $\epsilon$ = molar extinction coefficient (molar absorptivity)<br>$c$ = sample's concentration<br>$l$ = path length |

### Chapter 2: Atomic and Electronic Structure

|                                                             |                                               |                                                                                                                                                                                                            |
|-------------------------------------------------------------|-----------------------------------------------|------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| <b>Energy of a photon</b>                                   | $E_{\text{photon}} = hf = \frac{hc}{\lambda}$ | $h$ = Planck's constant<br>( $6.63 \times 10^{-34} \text{ J} \cdot \text{s}$ )<br>$f$ = photon's frequency<br>$c$ = speed of light<br>( $3.0 \times 10^8 \text{ m/s}$ )<br>$\lambda$ = photon's wavelength |
| <b>Absorption/Emission Line Spectra</b>                     | $\Delta E = E_{\text{photon}}$                |                                                                                                                                                                                                            |
| <b>Kinetic Energy of an electron (Photoelectric Effect)</b> | $KE_{e^-} = E_{\text{photon}} - \phi$         | $\phi$ = work function (minimum energy needed to ionize electron)                                                                                                                                          |

### Chapter 7: Chemical Solutions

|                                                 |                                                                |                                                                                                                                                          |
|-------------------------------------------------|----------------------------------------------------------------|----------------------------------------------------------------------------------------------------------------------------------------------------------|
| <b>Molarity</b>                                 | $M = \frac{\text{moles}_{\text{solute}}}{L_{\text{solution}}}$ |                                                                                                                                                          |
| <b>Molality</b>                                 | $m = \frac{\text{moles}_{\text{solute}}}{kg_{\text{solvent}}}$ |                                                                                                                                                          |
| <b>Henry's Law</b>                              | $P_A = k_H[A]$                                                 | $P_A$ = partial pressure of gas A<br>$k_H$ = Henry's Law constant (varies per problem)<br>$[A]$ = conc. of gas A                                         |
| <b>Freezing Point Depression</b>                | $\Delta T_F = -iK_F m$                                         | $i$ = van't Hoff factor<br>$K_F$ = F.P. depression constant<br>$m$ = molality                                                                            |
| <b>Boiling Point Elevation</b>                  | $\Delta T_B = iK_B m$                                          | $i$ = van't Hoff factor<br>$K_B$ = B.P. depression constant<br>$m$ = molality                                                                            |
| <b>Vapor Pressure Depression (Raoult's Law)</b> | $P_{\text{solute}} = X_{\text{solute}} P_{\text{pure}}^0$      | $P_{\text{solute}}$ = VP of solution<br>$X_{\text{solute}}$ = mol fract of solvent<br>$P_{\text{pure}}^0$ = VP of pure solvent                           |
| <b>Osmotic Pressure (<math>\pi</math>)</b>      | $\pi = iMRT$                                                   | $M$ = molarity of solute<br>$i$ = van't Hoff factor<br>$R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$<br>$T$ = temp. in Kelvin |

### Chapter 5: Gases

|                                                |                                                                                |                                                                                                                             |
|------------------------------------------------|--------------------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------------------------------|
| <b>Pressure</b>                                | $P = \frac{F}{A}$                                                              | $F$ = force<br>$A$ = area                                                                                                   |
| <b>Average Kinetic Energy</b>                  | $KE_{\text{avg}} = \frac{3}{2} RT$                                             | $R = 8.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$                                                                      |
| <b>Root-Mean-Square Speed (<math>v</math>)</b> | $v = \sqrt{\frac{3RT}{M_m}}$                                                   | $R = 8.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$<br>$M_m$ = molar mass                                                |
| <b>Ideal Gas Law</b>                           | $PV = nRT$                                                                     | $n$ = # of moles<br>$R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$                                |
| <b>Boyle's Law</b>                             | $V \propto \frac{1}{P}$                                                        |                                                                                                                             |
| <b>Charles' Law</b>                            | $V \propto T$                                                                  |                                                                                                                             |
| <b>Avogadro's Law</b>                          | $V \propto n$                                                                  |                                                                                                                             |
| <b>Combined Gas Law</b>                        | $\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$                                |                                                                                                                             |
| <b>Standard Temp. &amp; Pressure (STP)</b>     | $P = 1 \text{ atm}$<br>$T = 273 \text{ K}$                                     | *1 mol of gas = 22.4 L at STP                                                                                               |
| <b>Standard Conditions</b>                     | All aqueous species @ 1M<br>All gaseous species @ 1 atm<br>$T = 298 \text{ K}$ |                                                                                                                             |
| <b>Density</b>                                 | $\frac{P(MM)}{RT} = \frac{m}{V}$                                               | $MM$ = molar mass<br>$R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$<br>$m$ = mass<br>$V$ = volume |

|                                          |                                                  |                                                                                                                 |
|------------------------------------------|--------------------------------------------------|-----------------------------------------------------------------------------------------------------------------|
| <b>Dalton's Law of Partial Pressures</b> | $P_{\text{total}} = P_A + P_B + \dots$           |                                                                                                                 |
| <b>Dalton's Law of Partial Pressures</b> | $P_A = X_A P_{\text{total}}$                     | $X_A$ = mol fraction of gas A                                                                                   |
| <b>Graham's Law of Effusion</b>          | $\frac{r_1}{r_2} = \sqrt{\frac{M_{m2}}{M_{m1}}}$ | $r$ = rate of effusion<br>$M$ = molar mass                                                                      |
| <b>Real Gas Equation</b>                 | $(P + \frac{an^2}{V^2})(V - nb) = nRT$           | $a$ & $b$ = constants specific to each gas<br>$\frac{an^2}{V^2}$ corrects for IMFs<br>$-nb$ corrects for volume |

### Chapter 8: Chemical Kinetics

|                           |                                                                                                                                                                                |                                                                                                                                                           |
|---------------------------|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------------------------------------------------------------|
| <b>General</b>            | $A + B \rightarrow C + D$                                                                                                                                                      | $k$ = rate constant<br>$m$ & $n$ = determined experimentally                                                                                              |
| <b>Rate Law</b>           | $\text{rate} = k[A]^m[B]^n$                                                                                                                                                    | $k$ = rate constant<br>$M$ = molarity<br>$s$ = seconds                                                                                                    |
| <b>Rate Constant</b>      | 0 order: $k = M^1 \cdot s^{-1}$<br>1 <sup>st</sup> order: $k = s^{-1}$<br>2 <sup>nd</sup> order: $k = M^{-1} \cdot s^{-1}$<br>3 <sup>rd</sup> order: $k = M^{-2} \cdot s^{-1}$ |                                                                                                                                                           |
| <b>Units</b>              |                                                                                                                                                                                |                                                                                                                                                           |
| <b>Arrhenius Equation</b> | $k = Ae^{-E_a/RT}$                                                                                                                                                             | $k$ = rate constant<br>$A$ = unique to each rxn<br>$E_a$ = act. energy<br>$R = 8.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$<br>$T$ = temp. in Kelvin |

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