#### Chapter 5

#### Section 5.1: Thermodynamics

- 1. Predict the spontaneity of a reaction based on the  $\Delta G$  value.
- G = free energy = amount of energy available to do work
- 6 G= exergonic; gives off energy; does not require energy input = spontaneous
- + 

  G=endergonic: absorb energy: requires energy input = nonspontaneous
- 2. Draw reaction energy diagrams for exergonic and endergonic reactions.
- 3. Calculate the energy content in foods from its nutrient molecules.

Nutrient Molecule	Example	Energy (Cal/g)	(kJ/z)
Carbohydrate	Table sugar, potatoes, flour	4 4	17
Protein	Meats, fish, beans	4	17
Fat	Oil, butter	9	38

- "Cal/g for each molecule has only 1 sig fig so the answer is rounded to 1 sig fig"."
  - Integrating chemistry: Low-Calorie Foods, Discovering the Concepts: Reaction Energy Diagrams.

Food that is high in fat are high in calories-more energy

- ""refer to numbers 2 & 3""
- 5. Determine how much activation energy there is
- Activation Energy energy necessary to align the reactant molecules and to cause them to collide with enough energy to form products
- \*For a reaction to begin reactants have to collide with enough energy/correct orientation\*
- 6. Types of energy (define and examples/symbols) (activation, entropy, free energy, enthalpy) what do they measure?
  - Enthalpy: represented as AH
  - Entropy: represented as AS
  - Free Energy: the amount of energy in a molecule that is available to do work, represented by the equation  $\triangle G = \triangle H T(\triangle S)$

Activation Energy: the energy necessary to align the reactant molecules and to cause them to collide with enough energy to form products, a property of a chemical reaction, can think of it as "the energy hill that molecules must overcome before they change into products"

# **Acs Thermodynamics Exam Study Guide**

Axel E. W. Müller

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