

1. According to the Bronsted—Lowry definition, which chemical species can function *both* as an acid and as a base?

- a. Cl^-
- b. SO_4^{2-}
- c. NH_4^+
- d. HCO_3^-
- e. H_3O^+

2. In the reaction $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$ which is an acid—base conjugate pair?

- a. H_2O and HCN
- b. H_2O and OH^-
- c. CN^- and H_2O
- d. HCN and OH^-

3. Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX ?

- a. It is less basic than a 1 M solution of NaY .
- b. It is more basic than a 1 M solution of NaY .
- c. It yields a neutral solution.
- d. It is more concentrated than a 1 M solution of NaY .

4. The pH of a 0.03 M HCl solution is

- a. 1.5
- b. 2.5
- c. 3.5
- d. 12.5

5. The pH of a 1.0×10^{-2} M $\text{Ba}(\text{OH})_2$ solution at 25°C is

- a. 2.7
- b. 3.0
- c. 11.0
- d. 11.3

6. Which is the strongest acid?

- a. HClO
- b. HClO_2
- c. HClO_3
- d. HClO_4

7. Which salt reacts with water (hydrolyzes) to produce a basic solution?

- a. $\text{NaC}_2\text{H}_3\text{O}_2$
- b. NaNO_3
- c. NH_4Cl
- d. BaSO_4

8. The *weakest* of the bases listed is

Acid	Conjugate Base	K_a (Ionization Constant of Acid)
HCl	Cl^-	100% ionized
HSO_4^-	SO_4^{2-}	1.2×10^{-2}
H_2S	HS^-	5.7×10^{-8}
HS^-	S^{2-}	1.2×10^{-13}

- a. Cl^-
- b. CN^-
- c. HS^-
- d. S^{2-}
- e. SO_4^{2-}

9. The oxide of which element will react with water to form the strongest acid?

	Main Groups							
Period	I	II	III	IV	V	VI	VII	O
First								
Second	X	Y	Z	P		Q	S	U
Third	W					R	T	M

- a. W
- b. M
- c. P
- d. R
- e. Z

10. Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate $\text{KC}_2\text{H}_3\text{O}_2$ is less alkaline than a 0.10 M solution of potassium cyanide, KCN ?

- a. Hydrocyanic acid is a weaker acid than acetic acid.
- b. Hydrocyanic acid is less soluble in water than acetic acid.
- c. Cyanides are less soluble than acetates.
- d. Acetic acid is a weaker acid than hydrocyanic acid.

11. In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH , the properties of the solution at the equivalence point will correspond exactly to the properties of

- a. a 0.100 M sodium solution.
- b. a 0.0500 M sodium hydroxide solution.
- c. a 0.0500 M benzoic acid solution.
- d. a 0.0500 M sodium benzoate solution.

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A. Wallace Hayes



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