

ACS Chemistry General Exam, (Answered) Complete Verified Solution 2023/2024

Layout of an element (mass # & atomic #)

atomic number =

of protons

mass number =

of protons + neutrons

the number of protons in a neutral atom will equal the

of electrons

how to find the correct number of neutrons

mass number - atomic number

Forms of the same elements that differ in amount of neutrons

isotopes

What particle, if lost from the nucleus, would NOT cause a change in the atomic number?

a neutron because it has no charge

When alpha particles were shot at a metal foil target, most passed through without deflections, while others deflected at large angles. What did this suggest to Rutherford?

that the atoms of the metal were mostly empty space, while the nucleus consisted of most of the mass which included highly condensed positive particles (which caused the deflection).

An atom has a valence shell configuration of $1s^1$. To which group of elements on the periodic table does it belong?

1A

alkali metals

electromagnetic spectrum

The greater the wavelength of a photon..?

the lower its frequency

The greater the frequency of a photon...?

the shorter its wavelength

Light energy formula

What does C in the light energy formula represent?

Give the ground-state electron configuration of an atom in the second period

$1s^2 2s^1$

Electrons configuration rule that states electrons do not pair until they have to. (put one electrons in each electron domain then pair up)

Hund's rule

An experimental phenomena associated with atoms having unpaired electrons - which states such substances are attracted to magnetic fields

Paramagnetism

Which drop in energy level results in the greatest emission of energy?

$n=2$ to $n=1$

quantum numbers

Acs General Chemistry Exam 2013

Annelies Wilder-Smith



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