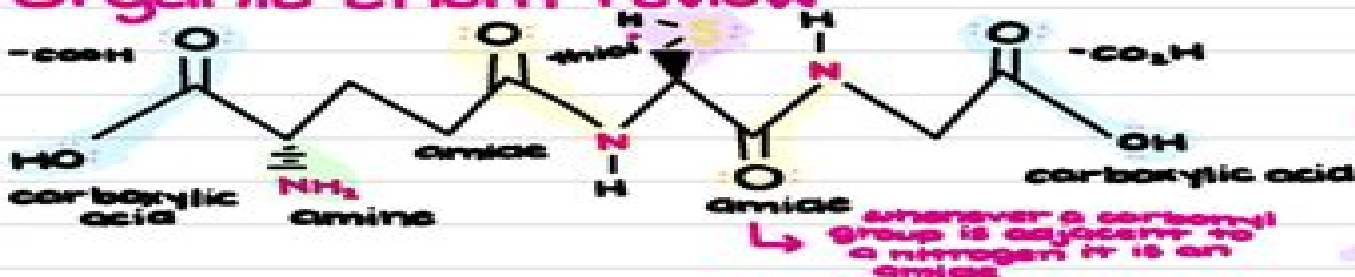


# Organic chem review



Draw Lewis structure:



notes:

- Hydrogen only forms one compound
- if carbon is neutral it will form 4 bonds
- if carbon has a charge it will form 3 bonds ( $-\text{C}^-$ )
- when oxygen is neutral it will typically form 2 bonds



Carbon has 4 bonds so there needs to be a triple bond

and one lone pair ( $\text{O:}$ )

when nitrogen is neutral it will usually form 3 bonds and a lone pair ( $-\text{N}-$ )



Key note:

$$\equiv > = > -$$

$\sigma > \pi$

sigma bond      pi bond

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**Didier Musso**



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