

ACS Practice Test 1

Acids & Bases:

- All are examples of Lewis acid–base reactions *except*
 - $\text{Cu}^{2+}_{(aq)} + 4\text{NH}_3_{(aq)} \rightleftharpoons [\text{Cu}(\text{NH}_3)_4]^{2+}_{(aq)}$
 - $\text{HCl}_{(g)} + \text{NH}_3_{(g)} \rightarrow \text{NH}_4\text{Cl}_{(s)}$
 - $\text{H}^+_{(aq)} + \text{OH}^-_{(aq)} \rightleftharpoons \text{H}_2\text{O}_{(l)}$
 - $2\text{Na}_{(s)} + \text{Cl}_{2(g)} \rightarrow 2\text{NaCl}_{(s)}$
- According to the Lewis definition, an acid is a species
 - having a hydrogen ion.
 - donating a pair of electrons.
 - accepting a pair of electrons.
 - accepting a hydrogen ion.
- All are potential Lewis bases *except*
 - NH_3
 - H_2O
 - CH_4
 - CN^-
- Which hydrated metal ion is most acidic under conditions of equal molar concentration in water?
 - Al^{3+}
 - Ba^{2+}
 - K^+
 - Zn^{2+}
 - Ag^+
- Which element in Group 5A forms the most basic oxide?
 - N
 - P
 - As
 - Sb
 - Bi

Atomic Structure:

- In what respect does an atom of magnesium differ from a magnesium ion (Mg^{2+})?
 - The ion has a more stable electronic arrangement than the atom.
 - The positive charge on the nucleus of the ion is two units greater than the nuclear charge on the atom.
 - The ion has two more protons than the atom.
 - The ion will react more readily with free Cl_2 .
 - The ion has two more planetary electrons than the atom.
- A certain atom has the symbol **Px**. What does this symbol tell about an atom of the element?
 - It has 148 electrons.
 - It has a valence of 4.
 - It has 221 protons in the nucleus.
 - It has 75 neutrons in the nucleus.
 - Its atomic mass is 73 u.
- The atomic mass of an element is 32.07 u and its atomic number is 16. The number of protons contained in the nucleus of an atom of this element is
 - 16
 - 31
 - 32
 - 48
- The atomic mass of an element is 32.07 u and its atomic number is 16. The element forms a simple ion. The ion will most likely have a charge of
 - 1–
 - 2–
 - 3+
 - 1+

Acs Chemistry Practice Test

**Kathleen M. Savage, Charity Anne
Dorgan**



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separated into first term and second term general chemistry material Each section contains 8 chapters of material that also aligns to most general chemistry textbooks for a seamless addition to study materials for students Each chapter is designed with an introductory section of the material including common representations and where to find this material in a textbook The second section provides worked examples of typical multiple choice questions including how the correct answer is determined as well as how the incorrect answers were determined Also included for each study problem is a listing of the corresponding practice questions that use that concept The final section is a series of practice problems to test the concepts collectively The key is provided on a separate page for all study and practice problems *Barron's AP Chemistry* Neil D. Jespersen, Pamela Kerrigan, 2016-03-01 Extensive test preparation for the AP Chemistry exam includes Six practice AP exams three diagnostic tests and three full length practice exams that are aligned with the upcoming new AP Chemistry exam All questions answered and explained A comprehensive subject review covering the structure of matter chemical bonding states of matter physical chemistry chemical reactions and all other test topics Study tips and test taking strategies This review book can be purchased alone or with an optional CD ROM that presents two additional full length practice tests with automatic scoring and fully explained answers **Professional Careers Sourcebook** Kathleen M. Savage, Charity Anne Dorgan, 1990 Provides a comprehensive overview of the literature and professional organizations that aid career planning and related research for 111 careers requiring college degrees or specialized education Validation, Verification, and Testing of Computer Software W. Richards Adrion, 1981

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