

5 • Reaction in Aqueous Solution

"ate" becomes "____ic acid"

"ite" becomes "____ous acid"

"ide" becomes "hydro____ic acid"

bromate	BrO_3^-	HBrO_3	bromic acid
periodate	IO_4^-	HIO_4	periodic acid
carbonate	CO_3^{2-}	H_2CO_3	carbonic acid
peroxide*	O_2^{2-}	H_2O_2	hydrogen peroxide
chloride	Cl^-	HCl	hydro-chloric acid
chlorite	ClO_2^-	HClO_2	chlorous acid
thiosulfate	$\text{S}_2\text{O}_3^{2-}$	$\text{H}_2\text{S}_2\text{O}_3$	thiosulfuric acid
sulfide	S^{2-}	H_2S	hydro-sulfuric acid
dichromate	$\text{Cr}_2\text{O}_7^{2-}$	$\text{H}_2\text{Cr}_2\text{O}_7$	dichromic acid
hypobromite	BrO^-	HBrO	hypobromous acid
sulfite	SO_3^{2-}	H_2SO_3	sulfurous acid
chromate	CrO_4^{2-}	H_2CrO_4	chromic acid
permanganate	MnO_4^-	HMnO_4	permanganic acid
iodate	IO_3^-	HIO_3	iodic acid
perbromate	BrO_4^-	HBrO_4	perbromic acid
cyanide	CN^-	HCN	hydro-cyanic acid
chlorate	ClO_3^-	HClO_3	chloric acid
nitrate	NO_3^-	HNO_3	nitric acid

* = be careful

ANSWERS to NAMING ACIDS

in sulfur compounds, add "ur"

in phosphorus compounds, add "or"

perchlorate	ClO_4^-	HClO_4	perchloric acid
bisulfate*	HSO_4^-	H_2SO_4	sulfuric acid
hypoiodite	IO^-	HIO	hypoiodous acid
bicarbonate*	HCO_3^-	H_2CO_3	carbonic acid
sulfate	SO_4^{2-}	H_2SO_4	sulfuric acid
iodite	IO_2^-	HIO_2	iodous acid
acetate	$\text{C}_2\text{H}_3\text{O}_2^-$	$\text{HC}_2\text{H}_3\text{O}_2$	acetic acid
iodide	I^-	HI	hydroiodic acid
bromide	Br^-	HBr	hydrobromic acid
hydroxide*	OH^-	HOH	water
phosphate	PO_4^{3-}	H_3PO_4	phosphoric acid
hypochlorite	ClO^-	HClO	hypochlorous acid
phosphite	PO_3^{3-}	H_3PO_3	phosphorous acid
oxide*	O^{2-}	H_2O	water
fluoride	F^-	HF	hydrofluoric acid
thiocyanate	SCN^-	HSCN	thiocyanic acid
bromite	BrO_2^-	HBrO_2	bromous acid
nitrite	NO_2^-	HNO_2	nitrous acid

Acids And Bases Worksheet Answers South Pasadena

Kristi Lew



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Table of Contents Acids And Bases Worksheet Answers South Pasadena

1. Understanding the eBook Acids And Bases Worksheet Answers South Pasadena
 - The Rise of Digital Reading Acids And Bases Worksheet Answers South Pasadena
 - Advantages of eBooks Over Traditional Books
2. Identifying Acids And Bases Worksheet Answers South Pasadena
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acids And Bases Worksheet Answers South Pasadena
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acids And Bases Worksheet Answers South Pasadena
 - Personalized Recommendations
 - Acids And Bases Worksheet Answers South Pasadena User Reviews and Ratings
 - Acids And Bases Worksheet Answers South Pasadena and Bestseller Lists
5. Accessing Acids And Bases Worksheet Answers South Pasadena Free and Paid eBooks

- Acids And Bases Worksheet Answers South Pasadena Public Domain eBooks
 - Acids And Bases Worksheet Answers South Pasadena eBook Subscription Services
 - Acids And Bases Worksheet Answers South Pasadena Budget-Friendly Options
6. Navigating Acids And Bases Worksheet Answers South Pasadena eBook Formats
- ePub, PDF, MOBI, and More
 - Acids And Bases Worksheet Answers South Pasadena Compatibility with Devices
 - Acids And Bases Worksheet Answers South Pasadena Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Acids And Bases Worksheet Answers South Pasadena
 - Highlighting and Note-Taking Acids And Bases Worksheet Answers South Pasadena
 - Interactive Elements Acids And Bases Worksheet Answers South Pasadena
8. Staying Engaged with Acids And Bases Worksheet Answers South Pasadena
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acids And Bases Worksheet Answers South Pasadena
9. Balancing eBooks and Physical Books Acids And Bases Worksheet Answers South Pasadena
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acids And Bases Worksheet Answers South Pasadena
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Acids And Bases Worksheet Answers South Pasadena
- Setting Reading Goals Acids And Bases Worksheet Answers South Pasadena
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acids And Bases Worksheet Answers South Pasadena
- Fact-Checking eBook Content of Acids And Bases Worksheet Answers South Pasadena
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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