

# Acids and Bases Worksheet 1

1. Finish the following reactions when acids are added to water.



2. What are the conjugate bases of these acids?

original acid	conjugate base
$\text{HNO}_3$	
$\text{H}_2\text{O}$	
$\text{H}_3\text{O}^+$	
$\text{H}_2\text{SO}_4$	
$\text{HBr}$	
$\text{HCO}_3^-$	

3. What are the conjugate acids of these bases?

original base	conjugate acid
$\text{OH}^-$	
$\text{H}_2\text{O}$	
$\text{HCO}_3^-$	
$\text{SO}_4^{2-}$	
$\text{ClO}_4^-$	

4. Which of the following represent conjugate acid-base pairs?



5. Calculate the  $[\text{H}^+]$  in a solution in which  $[\text{OH}^-] = 2.0 \times 10^{-2} \text{ M}$ . Is this solution acidic, neutral, or basic?

6. Calculate the  $[\text{OH}^-]$  in a solution in which  $[\text{H}^+] = 3.99 \times 10^{-5} \text{ M}$ . Is this solution acidic, neutral, or basic?

# Acids And Bases Worksheet 1 Answers

**S K Gupta**



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