

Unit 8 - Acids and Bases

Bronsted-Lowry Theory

The Bronsted-Lowry Theory for acids and bases states that:

- A **Bronsted-Lowry acid** is a H^+ donor
- A **Bronsted-Lowry base** is a H^+ acceptor



- In this example, NH_3 is a Bronsted-Lowry Base, because it accepts an H^+ ion to become NH_4^+ , while H_2O is a Bronsted-Lowry Acid because it donates an H^+ ion to NH_3 to become OH^-

Water and Amphiprotic species

Amphiprotic Species a compound that has the ability to **act as both a Bronsted-Lowry acid and base**

- This means it can donate (lose) an H^+ ion OR accept (gain) an H^+ ion

Water is an important example of this

- It can donate (lose) an H^+ ion to become OH^- (behaves like an BL acid)
- It can accept (gain) an H^+ ion to become H_3O^+ (behaves like a BL base)

Other examples of Amphiprotic Species include HCO_3^- and HSO_4^-

Types of Acids

Acids can be classified into certain categories based on the number of hydrogen ions/protons that they can donate in a given acid-base reaction.

The categories of acids are

- **Monoprotic:** can only donate one H^+ ion: eg. HCl , HNO_3
- **Diprotic:** can donate two H^+ ions: eg. H_2SO_4 , H_2CO_3
- **Triprotic:** can donate three H^+ ions: e.g. H_3PO_4

All carboxylic acids are monoprotic. Despite having more than one H^+ ion, they can only donate one.

Conjugate Acid-Base Pairs

- In order to find the conjugate base of an acid, we must remove a H^+ ion to the compound.
- In order to find the conjugate acid of a base, we must add a H^+ ion to the compound.
 - In general, Conjugate Acid-Base Pairs differ by one H^+ ion

Acids And Bases Unit Study Guide

David R. Klein



Acids And Bases Unit Study Guide:

Chemical Principles Student's Study Guide & Solutions Manual John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text *Pediatric Critical Care Study Guide* Steven E. Lucking, Frank A. Maffei, Robert F.

Tamburro, Neal J. Thomas, 2012-04-10 This is the first comprehensive study guide covering all aspects of pediatric critical care medicine It fills a void that exists in learning resources currently available to pediatric critical care practitioners The major textbooks are excellent references but do not allow concise reading on specific topics and are not intended to act as both text and study guide There are also several handbooks available but these are usually written for general pediatric residents and lack the advanced physiology and pathophysiology required for the higher level pediatric critical care practitioner **Organic Chemistry, 5e Student Study Guide and Solutions Manual** David R. Klein, Laurie S.

Starkey, 2025-03-18 Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems With Organic Chemistry Student Study Guide and Solutions Manual 5th Edition students can learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry *MCAT Complete; Medical College Admission Test Comprehensive Study Guide Physical Sciences; Biological Sciences; Verbal Reasoning* James Adams, Maryn Munley, Nathan Eisenbeis, Caroline Miller, Jesse Shetler, 2021-10-12 The MCAT The Medical College Admission Test MCAT is a standardized multiple choice examination designed to assess the examinee s problem solving critical thinking and knowledge of science concepts and principles prerequisite to the study of medicine Scores are reported in Physical Sciences Verbal Reasoning and Biological Sciences Study guide covers AAMC Association of American Medical Colleges content Physical Sciences Biological Sciences Verbal Reasoning Mathematics Concepts The Cell Chromosomes Reproduction Implantation Microorganisms Biochemistry Human Physiology The Heart The Lymphatic System GI Tract Musculoskeletal System Kidney Hormones Nerves Skin Genetics Populations and Evolution Elements Hund s Rule and Radiation The Periodic Table Covalent Bonds Molecular Shapes General Chemistry Kinetic Molecular Theory Phase Change Solutions Oxidation Numbers Entropy Acids and Bases Galvanic and Electrolytic Cells Carbon Stereochemistry Alkanes and Alkenes Hydrogen Bonding Alcohols Phenols Aldehydes and Ketones m Carboxylic Acids Ether Ammonia Amino Acids Carbohydrates Spectroscopy Separation and Distillation Vectors and Simple Motion Forces Circular and Projectile Motion Statics Center of Gravity Work and Energy Power and Momentum Stress and Strain Elasticity and Density Hydrostatic Pressure Fluids in Motion Electricity and Magnetism Plates Capacitors Voltage Batteries Resistors Magnetic Fields Waves and Periodic Motion Sound Waves Doppler Effect Simple Harmonic Motion Optics Mirrors 205 pages Organic Chemistry, 4e Student Solution Manual and Study

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