

1. In the list below, circle the properties of acids and box the properties of bases.

|                   |                                  |                         |
|-------------------|----------------------------------|-------------------------|
| Sour taste        | Feel slippery                    | Contains hydroxide ions |
| Turns litmus blue | Reacts with metals to form $H_2$ | Turns litmus red        |
| Bitter taste      | Contains hydrogen ions           | Corrosive               |

2. Complete the chart below:

| Acid-Base Definitions |                                  |                                   |
|-----------------------|----------------------------------|-----------------------------------|
| Type                  | Acid Definition                  | Base Definition                   |
| Arrhenius             | Donates an $H^+$ ion in solution | Donates an $OH^-$ ion in solution |
| Bronsted-Lowry        | Donates an $H^+$ ion             | Accepts an $H^+$ ion              |
| Lewis                 | Accepts an electron pair         | Donates an electron pair          |

3. What term(s) describe a substance that can act as an acid or a base? **Amphoteric or amphiprotic**
4. Neutralization reactions always produce a **SALT** and **WATER**
5. If the  $[H^+]$  is greater than the  $[OH^-]$ , the solution is acidic basic neutral
6. If the  $[OH^-]$  is greater than the  $[H^+]$ , the solution is acidic basic neutral
7. If the  $[H^+]$  is equal to the  $[OH^-]$ , the solution is acidic basic neutral
8. If the pH is 7, the solution is acidic basic neutral
9. If the pH is greater than 7, the solution is acidic basic neutral
10. If the pH is less than 7, the solution is acidic basic neutral
11. How many moles of lithium hydroxide are required to neutralize 19.2 grams of chloric acid?



$$(19.2 \text{ g } HClO_3) \times (1 \text{ mol } HClO_3 / 84 \text{ g } HClO_3) \times (1 \text{ mol } LiOH / 1 \text{ mol } HClO_3) = \underline{0.228 \text{ mol } LiOH}$$

12. How many grams of calcium hydroxide are required to neutralize 132.4 grams of nitric acid?



$$(132.4 \text{ g } HNO_3) \times (1 \text{ mol } HNO_3 / 63 \text{ g } HNO_3) \times (1 \text{ mol } Ca(OH)_2 / 2 \text{ mol } HNO_3) \times (74 \text{ g } Ca(OH)_2 / 1 \text{ mol } Ca(OH)_2) = \underline{77.76 \text{ g } Ca(OH)_2}$$

# Acids And Bases Chapter Review Chemistry Answers

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## **Table of Contents Acids And Bases Chapter Review Chemistry Answers**

1. Understanding the eBook Acids And Bases Chapter Review Chemistry Answers
  - The Rise of Digital Reading Acids And Bases Chapter Review Chemistry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Acids And Bases Chapter Review Chemistry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acids And Bases Chapter Review Chemistry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Acids And Bases Chapter Review Chemistry Answers
  - Personalized Recommendations
  - Acids And Bases Chapter Review Chemistry Answers User Reviews and Ratings
  - Acids And Bases Chapter Review Chemistry Answers and Bestseller Lists

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  - Acids And Bases Chapter Review Chemistry Answers Public Domain eBooks
  - Acids And Bases Chapter Review Chemistry Answers eBook Subscription Services
  - Acids And Bases Chapter Review Chemistry Answers Budget-Friendly Options
6. Navigating Acids And Bases Chapter Review Chemistry Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Acids And Bases Chapter Review Chemistry Answers Compatibility with Devices
  - Acids And Bases Chapter Review Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Acids And Bases Chapter Review Chemistry Answers
  - Highlighting and Note-Taking Acids And Bases Chapter Review Chemistry Answers
  - Interactive Elements Acids And Bases Chapter Review Chemistry Answers
8. Staying Engaged with Acids And Bases Chapter Review Chemistry Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Acids And Bases Chapter Review Chemistry Answers
9. Balancing eBooks and Physical Books Acids And Bases Chapter Review Chemistry Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Acids And Bases Chapter Review Chemistry Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Acids And Bases Chapter Review Chemistry Answers
  - Setting Reading Goals Acids And Bases Chapter Review Chemistry Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acids And Bases Chapter Review Chemistry Answers
  - Fact-Checking eBook Content of Acids And Bases Chapter Review Chemistry Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning



- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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