

ACIDS, BASES, & THE pH SCALE

Name: **Answer Key**

Date: _____

Pd: _____

PART 1: ACIDS AND BASES

Comparing Acids and Bases	
Acids	Bases (Alkalines)
<ul style="list-style-type: none">• Sour taste, leaves burning sensation• Increase hydrogen ions (H^+) in solution• Example: HCl $H^+ + Cl^-$• Turns litmus (pH) paper to red	<ul style="list-style-type: none">• Sharp bitter taste, feels slippery to touch• Decrease hydrogen ions (H^+) in solution (Increase hydroxide ions (OH^-) in solution)• Example: NaOH $Na^+ + OH^-$• Turns litmus (pH) paper to blue

PART 2: THE pH SCALE

- The **pH Scale** is a **measurement** of how acidic or how basic a solution is.
- The scale goes from **0** to **14**.
- **Compounds** are **acidic** if they have a pH **lower** than **7**.
- Compounds with a pH **higher** than **7** are **basic** or **alkaline**.
- Compounds with a pH **of 7**, are **neutral**.

0 1 2 3 4 5 6 7 8 9 10 11 12 13 14

Increased concentration of **H^+** ions
Decreased concentration of **OH^-** ions

Decreased concentration of **H^+** ions
Increased concentration of **OH^-** ions

pH below 7 = **acids**

pH at 7 = **neutral**
(**pure water**)

pH above 7 = **bases**

PART 3: BUFFERS

- A **buffer** is a **weak acid** or **base** that **prevents** sudden **changes** in pH.
 - Maintains **homeostasis**
- A buffer works by **binding** to H^+ ions **releasing** H^+ ions when the solution pH changes.
- Buffers **bind** to H^+ ions when the **concentration** of H^+ ions in solution **increases**.
- **NaOH** **$Na^+ + OH^-$** (The OH^- will bind to H^+ ions to form H_2O – **neutralization**)
- Buffers **release** H^+ ions when the **concentration** of H^+ ions in solution **decreases**.
HCl **$H^+ + Cl^-$**

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VM Jensen



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Acid-base Behavior in Aprotic Organic

Solvents Marion Maclean Davis, 1968 A unified picture of acid base behavior in aprotic organic solvents is presented based on an extensive survey of the literature and experimental results of the author and associates Evidence given to support this picture includes data pertaining to colligative properties of acids bases and salts and also conductance dielectric constants distribution between immiscible solvents and spectral absorption in the infrared visible and ultraviolet The acids upon which attention is centered are proton donor compounds that are measurably ionized in water such as aliphatic and aromatic carboxylic acids substituted phenols and mineral acids The bases of principal interest are likewise compounds capable of forming ions in water for example aliphatic and aromatic amines and derivatives of guanidine or pyridine The solvents emphasized are hydrocarbons and halohydrocarbons but data for dipolar aprotic solvents for example acetone acetonitrile and nitrobenzene are included Contrasts in acid base behavior and in acidity and basicity scales in aprotic and water like solvents are discussed The role of hydrogen bonding in aprotic solvents is discussed at length Important types of hydrogen bonded structures include chelate rings self associated acids bases and salts hydrogen bonded ion pairs and homo and heteroconjugate cations and anions Examples are given in which hydrogen bonding of these types affects such properties as the absorption spectrum of a salt the catalytic effect of an acid and the accurate location of a titration endpoint Author

NBS Monograph United States. National Bureau of Standards, 1959 **Handbook of Solvents, Volume 1** George

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