

## Acid and Base Worksheet - Answers

- 1) Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair:
- a)  $\text{HNO}_3 + \text{OH}^- \rightarrow \text{H}_2\text{O} + \text{NO}_3^-$   
 **$\text{HNO}_3$  and  $\text{NO}_3^-$  make one pair  $\text{OH}^-$  and  $\text{H}_2\text{O}$  make the other**
- b)  $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{NH}_3^+ + \text{OH}^-$   
 **$\text{CH}_3\text{NH}_2$  and  $\text{CH}_3\text{NH}_3^+$  make one pair  $\text{OH}^-$  and  $\text{H}_2\text{O}$  make the other**
- c)  $\text{OH}^- + \text{HPO}_4^{2-} \rightarrow \text{H}_2\text{O} + \text{PO}_4^{3-}$   
 **$\text{HPO}_4^{2-}$  and  $\text{PO}_4^{3-}$  make one pair  $\text{OH}^-$  and  $\text{H}_2\text{O}$  make the other**
- 2) The compound NaOH is a base by all three of the theories we discussed in class. However, each of the three theories describes what a base is in different terms. Use your knowledge of these three theories to describe NaOH as an Arrhenius base, a Brønsted-Lowry base, and a Lewis base.
- **NaOH is an Arrhenius base because it creates  $\text{OH}^-$  ions when placed in water.**
  - **NaOH is a Brønsted-Lowry base because it accepts  $\text{H}^+$  ions from acids.**
  - **NaOH is a Lewis base because the lone pairs on the hydroxide ion can be donated to other compounds.**
- 3) When hydrogen chloride reacts with ammonia, ammonium chloride is formed. Write the equation for this process, and indicate which of the reagents is the Lewis acid and which is the Lewis base.
- $\text{HCl} + \text{NH}_3 \rightarrow \text{NH}_4\text{Cl}$   
 **$\text{NH}_3$  is a Lewis base because it uses its lone pair electrons to pull a hydrogen atom from hydrochloric acid.**  
 **$\text{HCl}$  is a Lewis acid because it accepts electrons from  $\text{NH}_3$  when the H is transferred.**
- 4) Write an equation for each of the following reactions and indicate what each reaction produces: (For each of the following oxides, indicate if it is an acid or basic anhydride).
- a. Carbon dioxide and water  
 $\text{CO}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3$   
**Acid anhydride**
- b. Sodium oxide and water  
 $\text{Na}_2\text{O} + \text{H}_2\text{O} \rightarrow 2\text{NaOH(aq)}$   
**Basic anhydride**
- c. Dichlorine trioxide and water  
 $\text{Cl}_2\text{O}_3 + \text{H}_2\text{O} \rightarrow 2\text{HClO}_2$   
**Acidic anhydride**
- d. Beryllium oxide and water  
 $\text{BeO} + \text{H}_2\text{O} \rightarrow \text{Be(OH)}_2$   
**Basic anhydride**

# Acids Bases Chemistry Test Answers

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