

# Acids and Bases

| Definitions       | Acid                    | Base                    |
|-------------------|-------------------------|-------------------------|
| 1. Bronsted-Lowry | H <sup>+</sup> donor    | H <sup>+</sup> acceptor |
| 2. Lewis          | e <sup>-</sup> acceptor | e <sup>-</sup> donor    |

## Typical Acid/Base Reaction:



**K<sub>eq</sub> for this example:** **K<sub>a</sub> for this example:**

$$K_{eq} = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{H}_2\text{O}][\text{HA}]} \quad K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$$

## In general:

- Strong acid → Weak conj. base
- Weak acid → Strong conj. base
- Strong base → Weak conj. acid
- Weak base → Strong conj. acid

## Predicting equilibria:

Start by **labeling the pK<sub>a</sub>s** in the reaction (see list)

Reaction Arrows:

- R<sub>2</sub>N **arrows** point towards the side with the **weaker acid** (higher pK<sub>a</sub>)

Is it reversible or irreversible?

- Look at the **difference of pK<sub>a</sub>s** between reactants/products
- $\Delta \text{pK}_a > 8 \Rightarrow$  irreversible rxn
- $\Delta \text{pK}_a < 8 \Rightarrow$  reversible rxn

**Example:** Predict equilibrium using pK<sub>a</sub>s, is it reversible or irreversible?



**Acidities:**

1. Label pK<sub>a</sub>s
2. Know arrows will point towards **weaker acid** (higher pK<sub>a</sub>), so left
3. Is it reversible or irreversible? (25-16 > 8), so irreversible
4. Final arrows:  $\rightarrow$

Things to know about acid/base reactions:

**Equilibrium constant (K<sub>eq</sub>):** a ratio that tells the position of a chemical equilibrium

$$K_{eq} = \frac{[\text{products}]}{[\text{reactants}]}$$

**Trends:**

- $K_{eq} > 1 \Rightarrow$  Favors products
- $K_{eq} = 1 \Rightarrow$  At equilibrium
- $K_{eq} < 1 \Rightarrow$  Favors reactants (non doesn't work)

**Acid Dissociation Constant (K<sub>a</sub>):** measure of acid strength

$$K_a = K_{eq} [\text{H}_2\text{O}]$$

**Trends:**

- $\uparrow K_a \Rightarrow$ 
  - Strong acid
  - Full dissociation
  - High [products] (large numerator)
- $\downarrow K_a \Rightarrow$ 
  - Weak acid
  - Partial/incomplete dissociation
  - High [reactants] (large denominator)

$$\text{pK}_a = -\log K_a$$

**Trends:**

- $\downarrow \text{pK}_a =$  Strong acid  $\uparrow \text{pK}_a =$  Weak acid
- $\uparrow \text{pK}_a =$  Weak acid  $\downarrow \text{pK}_a =$ 
  - $\text{pK}_a < 1 \Rightarrow$  Strong acid
  - $\text{pK}_a 1-3 \Rightarrow$  Moderately strong acid
  - $\text{pK}_a 3-5 \Rightarrow$  Weak acid
  - $\text{pK}_a > 5 \Rightarrow$  Very weak acid

## pK<sub>a</sub>s to Know

- sp<sup>3</sup> C-H 51
- sp<sup>2</sup> C-H 44
- sp C-H 25
- Water H<sub>2</sub>O = 16 (15.7)
- Ethanol OH = 16 (15.9)
- Ammonia NH<sub>3</sub> 38
- Acetic Acid CO<sub>2</sub>H = 5 (4.76)
- HCl -7

# Acids And Bases An Introduction Study Guide

**Christian Drosten**



## **Acids And Bases An Introduction Study Guide:**

## **Acids And Bases An Introduction Study Guide** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Acids And Bases An Introduction Study Guide**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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### **Table of Contents Acids And Bases An Introduction Study Guide**

1. Understanding the eBook Acids And Bases An Introduction Study Guide
  - The Rise of Digital Reading Acids And Bases An Introduction Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Acids And Bases An Introduction Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acids And Bases An Introduction Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Acids And Bases An Introduction Study Guide
  - Personalized Recommendations
  - Acids And Bases An Introduction Study Guide User Reviews and Ratings
  - Acids And Bases An Introduction Study Guide and Bestseller Lists

5. Accessing Acids And Bases An Introduction Study Guide Free and Paid eBooks
  - Acids And Bases An Introduction Study Guide Public Domain eBooks
  - Acids And Bases An Introduction Study Guide eBook Subscription Services
  - Acids And Bases An Introduction Study Guide Budget-Friendly Options
6. Navigating Acids And Bases An Introduction Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Acids And Bases An Introduction Study Guide Compatibility with Devices
  - Acids And Bases An Introduction Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Acids And Bases An Introduction Study Guide
  - Highlighting and Note-Taking Acids And Bases An Introduction Study Guide
  - Interactive Elements Acids And Bases An Introduction Study Guide
8. Staying Engaged with Acids And Bases An Introduction Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Acids And Bases An Introduction Study Guide
9. Balancing eBooks and Physical Books Acids And Bases An Introduction Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Acids And Bases An Introduction Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Acids And Bases An Introduction Study Guide
  - Setting Reading Goals Acids And Bases An Introduction Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acids And Bases An Introduction Study Guide
  - Fact-Checking eBook Content of Acids And Bases An Introduction Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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