

What are Acids, Bases, and Salts?

The Swedish chemist Svante Arrhenius introduced the theory of ionization and used this theory to explain much about the behavior of acids and bases.

- An Arrhenius acid is defined as any compound that dissociates in aqueous solution to form Hydrogen ions.



- An Arrhenius base is defined as any compound that dissociates in aqueous solution to form Hydroxide ions.



- Salts are compounds that dissociate in aqueous solution releasing neither Hydrogen ions nor Hydroxide ions.



Using the Arrhenius definition, classify the following examples as acids, bases, or salts:

HBr ACID

Mg(OH)₂ BASE

HCl ACID

KNO₃ SALT

HFO₄ ACID

Ba(OH)₂ BASE

KCl SALT

H₃PO₄ ACID

HClO ACID

Al(OH)₃ BASE

KC₂H₃O₂ SALT

NaCl SALT

Acid Base Or Salt Worksheet Answers Instructional Fair

Brian Knapp

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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