

Name: _____ Date: _____ Pd. : _____



Webquest: Acids and Bases

Directions:

- You may hand-write your answers on this page, or access this word document from Mrs. Leibacher's website and type your answers.
- You must submit this webquest EITHER in hard copy (written on the worksheet) or electronically (Google Docs, e-mail, memory stick, burned to CD, submitted on My Big Campus, etc.). Be sure to TELL ME how you are turning in your work!
- Write your answers in complete sentences.
- This assignment counts as a quiz grade (100 pts).

Chem4 Kids Website:

http://www.chem4kids.com/files/react_acidbase.html

1. What do scientists use to measure pH? Describe this tool
2. What is the difference between a strong acid/base and a weak acid/base?

VisionLearning Website:

http://www.visionlearning.com/library/module_viewer.php?mid=58

3. Describe the physical properties of Acids:
4. Describe the physical properties of Bases:
5. Give 3 examples of acids and 3 examples of bases from the pH Scale on the site:

Examples of Acids	Examples of Bases

Acid And Base Webquest Answers

H. L. Finston, Allen C. Rychtman



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the magnitude of the interaction energies of solvent molecules with especially ions the ability of solvents to participate in hydrogen bonding and to accept or donate electron pairs is seen to be crucial Experimental methods for determining dissociation constants are described in detail In the remaining chapters dissociation constants of a wide range of acids in three distinct classes of solvents are discussed protic solvents such as alcohols which are strong hydrogen bond donors basic polar aprotic solvents such as dimethylformamide and low basicity and low polarity solvents such as acetonitrile and tetrahydrofuran Dissociation constants of individual acids vary over more than 20 orders of magnitude among the solvents and there is a strong differentiation between the response of neutral and charged acids to solvent change Ion pairing and hydrogen bonding equilibria such as between phenol and phenoxide ions play an increasingly important role as the solvent polarity decreases and their influence on acid base equilibria and salt formation is described

Acid-base Chemistry

Clifford W. Hand, Harry Lyon Blewitt, 1986 Very Good No Highlights or Markup all pages are intact **Science**, 2017

Acid-base and Electrolyte Balance Gösta Rooth, 1979 *Acid-Base Diagrams* Heike Kahlert, Fritz Scholz, 2013-07-31

Understanding acid base equilibria made easy for students in chemistry biochemistry biology environmental and earth sciences Solving chemical problems be it in education or in real life often requires the understanding of the acid base equilibria behind them Based on many years of teaching experience Heike Kahlert and Fritz Scholz present a powerful tool to meet such challenges They provide a simple guide to the fundamentals and applications of acid base diagrams avoiding complex mathematics This textbook is richly illustrated and has full color throughout It offers learning features such as boxed results and a collection of formulae *Acid-base Indicators* Izaak Maurits Kolthoff, 1937 **Acid-Base Equilibria -**

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