

- 6) In the first trial, the rate constant, k , obtained at 25°C is indicated in a spreadsheet along with a calculated value of k . What is the rate constant?

$$\begin{array}{l}
 \text{Reaction: } 2\text{NO}_2 \rightarrow 2\text{NO} + \text{O}_2 \\
 \text{Rate: } -\frac{1}{2} \frac{d[\text{NO}_2]}{dt} = \frac{d[\text{NO}]}{dt} = \frac{d[\text{O}_2]}{dt} \\
 \text{Initial } [\text{NO}_2] = 0.0100 \text{ M} \\
 \text{Initial } [\text{NO}] = 0.0000 \text{ M} \\
 \text{Initial } [\text{O}_2] = 0.0000 \text{ M} \\
 \text{Final } [\text{NO}_2] = 0.0040 \text{ M} \\
 \text{Final } [\text{NO}] = 0.0060 \text{ M} \\
 \text{Final } [\text{O}_2] = 0.0030 \text{ M} \\
 \text{Time } t = 10.0 \text{ min} \\
 \text{Rate } k = \frac{1}{2} \frac{d[\text{NO}_2]}{dt} = \frac{1}{2} \frac{0.0100 - 0.0040}{10.0} = 3.0 \times 10^{-4} \text{ s}^{-1}
 \end{array}$$

- 7) Predict whether each change will shift the equilibrium left, right, or have no effect for the following equilibrium reaction:



- adding more CO_2 → CO_2 will shift CO_3^{2-}
- adding more H_2O → H_2O will shift CO_3^{2-}
- adding a catalyst to the reaction mixture → CO_2 will shift CO_3^{2-}
- removing H_2O → H_2O will shift CO_3^{2-}
- adding more CO_3^{2-} → CO_3^{2-} will shift CO_2

- 8) For each change and reaction, determine ΔH , and q .

- a. $\text{H}_2\text{O} \rightarrow \text{H}_2\text{O}(g)$

$$\begin{array}{l}
 \Delta H = \text{endothermic} \\
 q = \text{endothermic} \\
 \Delta H = \text{endothermic} \\
 q = \text{endothermic} \\
 \Delta H = \text{endothermic} \\
 q = \text{endothermic}
 \end{array}$$

- b. $\text{H}_2\text{O} \rightarrow \text{H}_2\text{O}(l)$

$$\begin{array}{l}
 \Delta H = \text{exothermic} \\
 q = \text{exothermic} \\
 \Delta H = \text{exothermic} \\
 q = \text{exothermic}
 \end{array}$$

- 9) Predict which system will be at equilibrium with the surroundings when it has reached equilibrium. $\Delta H = 10.0 \text{ kJ}$

- a. H_2O

- b. $\text{H}_2\text{O}(g)$

- c. $\text{H}_2\text{O}(l)$

- d. $\text{H}_2\text{O}(s)$

$$\begin{array}{l}
 \Delta H = 10.0 \text{ kJ} \\
 \Delta H = 10.0 \text{ kJ} \\
 \Delta H = 10.0 \text{ kJ} \\
 \Delta H = 10.0 \text{ kJ}
 \end{array}$$

- 10) Predict all the following results which, along with their appropriate reason, result from the following reaction. The reaction involves water, a gas, and $\text{H}_2\text{O}(g)$.



$\text{H}_2\text{O}(g)$ gas will shift $\text{H}_2\text{O}(l)$ when $\Delta H = 10.0 \text{ kJ}$

Acid And Base Reactions Study Guide Answers

**T. W. Graham Solomons, Craig B.
Fryhle**



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