

Solutions, acids, and bases

1. VOCABULARY REVIEW OF SOLUTIONS

Solute - what is **dissolved** in the solvent to form a solution.

Solvent - the substance that **dissolves** the solute, ordinarily a liquid when the solute and the solvent come together to form a **homogeneous mixture**.

Miscible - two mix in any proportions (have to be mixed) - liquids will pour each other

Immiscible - don't mix in any proportions (have to be mixed)

Examples of miscible/immiscible - ethanol/water/methanol, etc. oil/water/acetone

Solute - aspect of being **dissolved** in liquid

Insoluble - a substance (tends to be a solid) that **will not dissolve** in a solvent. An example is sand and water

Suspension - looks like a **solution** when **stirred**, but **settles** a bit of water (suspensions). Also it is a mixture - gaseous mixture where the small solid particles of a substance are spread throughout a liquid without settling.

Unsaturated - solutes with solubility chart when the temperature would be enough to make a substance dissolve but at a greater amount of solute solubility **less solute dissolves than can** (ex. salt @ 20°C temp)

Saturated - solutes with solubility chart when the temperature-solubility relationship is at the line. In other words a **solution with solute has reached** at the maximum possible amount.

Supersaturated - contains **more than the max. amount** of solute that can be dissolved at the temperature. In the solubility chart this would mean the grams of solute would go above the line.

2. SOLUBILITY VS. TEMPERATURE CHART



Above the line - **supersaturated**

At the line - **saturated**

Under the line - **unsaturated**

3. METHODS OF SEPARATION

Filtration - separate particles by **size**

Distillation - separate by **boiling point** - some mix liquids to be separated.

Evaporation - allows one to remove a **liquid from a solid** by **transforming** the liquid to a **gaseous state**.

Centrifugation - separating molecules by their densities by spinning them apart.

Decanting - process of separating insoluble or suspension solutions by **pouring** away the liquid from the solid.

Chromatography - separation by the **rates** at which components of different solutions move through a medium (ex. gel electrophoresis)

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