Acid and Base Worksheet - Answers

- Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair:
 - a) HNO₃ + OH¹ → H₂O + NO₃¹¹
 HNO₃ and NO₃¹ make one pair OH¹ and H₂O make the other
 - b) CH₂NH₂ + H₂O → CH₃NH₃* + OH⁴
 CH₃NH₂ and CH₃NH₃* make one pair OH⁴ and H₂O make the other
 - c) OH¹ + HPO₁² → H₂O + PO₁² HPO₁² and PO₁³ make one pair OH¹ and H₂O make the other
- 2) The compound NaOH is a base by all three of the theories we discussed in class. However, each of the three theories describes what a base is in different terms. Use your knowledge of these three theories to describe NaOH as an Arrhenius base, a Brønsted-Lowry base, and a Lewis base.
 - NaOH is an Arrhenius base because it creates OH⁻¹ ions when placed in water.
 - NaOH is a Brensted-Lowry base because it accepts H^{*1} ions from acids.
 - NaOH is a Lewis base because the lone pairs on the hydroxide ion can be donated to other compounds.
- When hydrogen chloride reacts with ammonia, ammonium chloride is formed. Write the equation for this process, and indicate which of the reagents is the Lewis acid and which is the Lewis base. HCI + NH₂ → NH₂CI

NH₂ is a Lewis base because it uses its lone pair electrons to pull a hydrogen atom from hydrochloric acid.

HCI is a Lewis acid because it accepts electrons from NH₂ when the H is

HCI is a Lewis acid because it accepts electrons from NH₃ when the H is transferred.

- 4) Write an equation for each of the following reactions and indicate what each reaction produces: (For each of the following oxides, indicate if it is an acid or basic anhydride).
 - Carbon dioxide and water
 CO₂ + H₂O → H₂CO₂
 Acid anhydride
 - b. Sodium oxide and water
 Na₂O + H₂O → 2NaOH(aq)
 Basic anhydride
 - Dichlorine trioxide and water
 Cl₂O₂ + H₂O → 2HClO₂
 Acidic anhydride
 - d. Beryllium oxide and water
 BeO + H₂O → Be(OH)₂
 Basic anhydride

Acid Base Worksheet Answers

Herb Strongin, Kara Strongin, Gloria Strongin

Acid Base Worksheet Answers:

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