

Acid-Base balance Questions and answers

Respiratory acidosis **CORRECT ANSWERS pH low, PaCO₂ high, K⁺ high,**

compensated by HCO₃ high to bring pH to normal

Metabolic acidosis **CORRECT ANSWERS pH low, HCO₃ low, K⁺ high**

compensated by PaCO₂ low

Respiratory alkalosis **CORRECT ANSWERS pH high, PaCO₂ low, K⁺ low,**

compensated by HCO₃ low to bring pH to normal

Metabolic alkalosis **CORRECT ANSWERS pH high, HCO₃ high, K⁺ low**

compensated by PaCO₂ high

Buffer system **CORRECT ANSWERS CO₂ + H₂O = H₂CO₃ = H(+) + HCO₃**

What causes

Metabolic alkalosis? **CORRECT ANSWERS loss of hydrogen ions or acids or**

an excess of base bicarbonate

For example:

- hypovolemia,**
- the loss of gastric fluid via NG suction or vomiting,**
- excessive bicarbonate intake,**
- massive transfusion of whole blood,**
- hyperaldosteronism.**

The nurse reviews the arterial blood gas results of a client and notes the following: pH 7.45, PaCO₂ of 30 mm Hg (30 mm Hg), and HCO₃⁻ of 20 mEq/L (20 mmol/L). The nurse analyzes these results as indicating which condition? **CORRECT ANSWERS Respiratory alkalosis, compensated**

Acid Base Question Packet

Patrick Regan



Acid Base Question Packet:

Green Chemistry Laboratory Manual for General Chemistry Sally A. Henrie, 2015-03-18 Green chemistry involves designing novel ways to create and synthesize products and implement processes that will eliminate or greatly reduce negative environmental impacts. Providing educational laboratory materials that challenge students with the customary topics found in a general chemistry laboratory manual, this lab manual enables students to see how green chemistry principles can be applied to real world issues. Following a consistent format, each lab experiment includes objectives, prelab questions, and detailed step by step procedures for performing the experiments. Additional questions encourage further research about how green chemistry principles compare with traditional more hazardous experimental methods. **CBSE Most Likely**

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