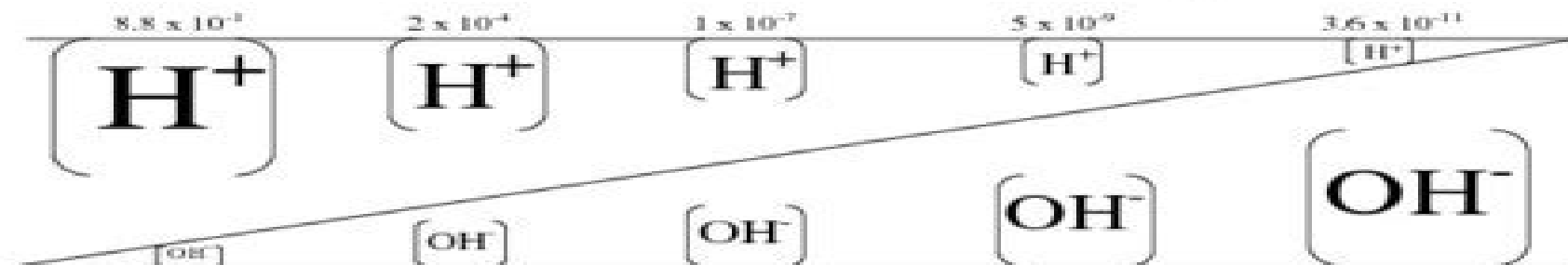


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pH PRACTICE



- Write the pH of each solution above the [H⁺]'s. $\text{pH} = -\log[\text{H}^+]$
- Label the "Z" diagram as "Acidic", "Basic" and "Neutral."
- Knowing that the $[\text{H}^+] \times [\text{OH}^-]$ always equals 1×10^{-14} , fill in the [OH⁻] for each of the five solutions in the "Z" Diagram. (1×10^{-14} is called the Dissociation Constant for water, K_w) ($1 \times 10^{-14} = 10 \times 10^{-15}$)
- Write the "pOH" of each solution below the [OH⁻]'s.
- $\text{pH} + \text{pOH}$ always equals _____.
- A solution of acid has $[\text{H}^+] = 3.0 \times 10^{-3} \text{ M}$
 - Calculate the [OH⁻] _____
 - Calculate the pH _____ the pOH _____
- A solution of base has an $[\text{OH}^-] = 4.25 \times 10^{-5} \text{ M}$
 - Calculate the [H⁺] _____
 - Calculate the pH _____ the pOH _____
- Calculate the pH's of the following solutions:

| | |
|-------------------------------------|------|
| $2.53 \times 10^{-2} \text{ M HCl}$ | pH = |
| $2.53 \times 10^{-4} \text{ M HCl}$ | pH = |
| $2.53 \times 10^{-5} \text{ M HCl}$ | pH = |

A pH with 3 significant figures is written with ____ numbers after the decimal place.

Acid Base Equilibria Ph Practice

Daniela Niemeyer



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