#### Acid - Base Titration

Introduction: In this lab, we are using a laboratory technique called Titrations, which can be used to determine concentrations of various chemical substances. In a titration, a standard solution with a known molarity is reacted with another solution such that the concentration of the second solution can be calculated from the results. In order to titrate the solutions, a buret is used to accurately display the volume which is later used in calculations. As well, the indicator Phenolphthalein is used to show when the neutralization occurs or the equivalence point is reached, which is when the number of moles of acid equals the number of moles in the base.

Objective/Purpose: The objective of this lab is to titrate a hydrochloric acid solution of unknown concentration with standardized 0.5M sodium hydroxide and to use the titration results to calculate the molarity of the hydrochloric acid.

**Hypothesis:** If an acid is titrated with a base, the acidic solutions pH will gradually decrease as more base is added until the solutions reach an equivalence point, at which the solution should present with a light pink color, due to the phenolphthalein dissolved in the solution. As well, by using our data gathered from the experiment, we can make calculations to identify the molarity of the hydrochloric acid.

#### Materials:

- Suction Bulb
- Volumetric Pipet (10 mL)
- Buret (50 mL)
- Burest stand and clamp
- Erlenmeyer flask (250 mL)
- Lab Apron
- Safety Goggles
- Standardized NaOH solution (approx. 0.5 M, known exactly)
- Unknown HCl solution
- Phenolphthalein indicator solution

Procedure: Refer to the lab handout

#### Data and Observations:

#### Quantitative:

Part 1: Unknown Solution 1

Molarity of NaOH = 0.5M	Trial 1	Trial 2	

**Robert De Levie** 

Aqueous Acid-base Equilibria and Titrations Robert De Levie, 1999 This book will give students a thorough grounding in pH and associated equilibria material absolutely fundamental to the understanding of many aspects of chemistry It is in addition a fresh and modern approach to a topic all too often taught in an out moded way This book uses new theoretical developments which have led to more generalized approaches to equilibrium problems these approaches are often simpler than the approximations which they replace Acid base problems are readily addressed in terms of the proton condition a convenient amalgam of the mass and charge constraints of the chemical system considered. The graphical approach of Bjerrum Hagg and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria Based on these concentrations the proton condition can usually be simplified often leading directly to the value of the pH In the description of acid base titrations a general master equation is developed It provides a continuous and complete description of the entire titration curve which can then be used for computer based comparison with experimental data Graphical estimates of the steepness of titration curves are also developed from which the practicality of a given titration can be anticipated Activity effects are described in detail including their effect on titration curves The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements which are subject to activity corrections and balance equations and spectroscopic pH measurements which are not Finally an entire chapter is devoted to what the pH meter measures and to the experimental and theoretical uncertainties involved Acid-base Titrations in Nonaqueous Solvents James Sherwood Fritz, 1973

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