

Section 1 Acids

Goals

- To describe acids
- To make the distinction between strong and weak acids.
- To show the changes that take place on the particle level when acids dissolve in water.
- To show how you can recognize strong and weak acids.

The most important skills to develop in this section are to be able to recognize acids from names or formulas.

Identification of Acids and Bases: Identify each of the following as an Arrhenius strong acid, an Arrhenius weak acid, an Arrhenius strong base, or an Arrhenius weak base.

- HNO_3 **strong acid**
- lithium hydroxide **strong base**
- K_2CO_3 **weak base**
- hydrofluoric acid **weak acid**

Section 2 Acid Nomenclature

- *Goal: To describe how to convert between names and chemical formulas for acids.*

This section adds acids to the list of compounds for which you should be able to convert between names and formulas.

Formulas for Acids: Write the chemical formulas that correspond to the names

- hydrofluoric acid: This name has the form of a binary acid, hydro(root)ic acid, so its formula is HF(aq) .
- phosphoric acid: This name has the form of an oxyacid, (root)ic acid, so it contains hydrogen, phosphorus, and oxygen. Phosphate is PO_4^{3-} , so phosphoric acid is H_3PO_4 .

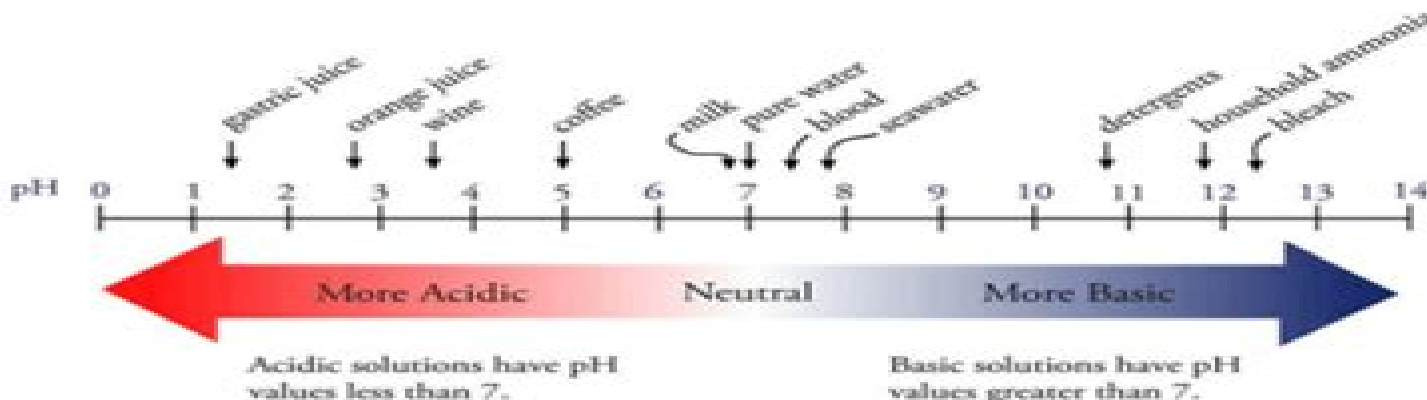
Naming Acids: Write the names that correspond to the chemical formulas

- HI : This is a binary acid, so its name has the form hydro(root)ic acid. HI(aq) is **hydroiodic acid**.
- $\text{HC}_2\text{H}_3\text{O}_2$: The name of the oxyanion $\text{C}_2\text{H}_3\text{O}_2^-$ is acetate, so $\text{HC}_2\text{H}_3\text{O}_2$ is **acetic acid**.

Section 3 Summary of Chemical Nomenclature

- *Goal: To review the process for converting between names and formulas for binary and ternary covalent, ionic compounds, acids, and oxyacids compounds.*

The ability to convert among names and formulas for chemical compounds is crucial to communication between chemists and chemistry students.



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Table of Contents Acid And Bases Study Guide Ch 15

1. Understanding the eBook Acid And Bases Study Guide Ch 15
 - The Rise of Digital Reading Acid And Bases Study Guide Ch 15
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid And Bases Study Guide Ch 15
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid And Bases Study Guide Ch 15
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid And Bases Study Guide Ch 15
 - Personalized Recommendations

- Acid And Bases Study Guide Ch 15 User Reviews and Ratings
- Acid And Bases Study Guide Ch 15 and Bestseller Lists
- 5. Accessing Acid And Bases Study Guide Ch 15 Free and Paid eBooks
 - Acid And Bases Study Guide Ch 15 Public Domain eBooks
 - Acid And Bases Study Guide Ch 15 eBook Subscription Services
 - Acid And Bases Study Guide Ch 15 Budget-Friendly Options
- 6. Navigating Acid And Bases Study Guide Ch 15 eBook Formats
 - ePub, PDF, MOBI, and More
 - Acid And Bases Study Guide Ch 15 Compatibility with Devices
 - Acid And Bases Study Guide Ch 15 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acid And Bases Study Guide Ch 15
 - Highlighting and Note-Taking Acid And Bases Study Guide Ch 15
 - Interactive Elements Acid And Bases Study Guide Ch 15
- 8. Staying Engaged with Acid And Bases Study Guide Ch 15
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid And Bases Study Guide Ch 15
- 9. Balancing eBooks and Physical Books Acid And Bases Study Guide Ch 15
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid And Bases Study Guide Ch 15
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acid And Bases Study Guide Ch 15
 - Setting Reading Goals Acid And Bases Study Guide Ch 15
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acid And Bases Study Guide Ch 15
 - Fact-Checking eBook Content of Acid And Bases Study Guide Ch 15

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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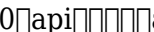

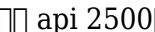




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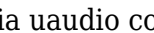
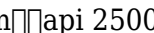
















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